CIVIL ENGINEERING

GATE | PSUs

ENVIRONMENTAL ENGINEERING

Volume - I : Study Material with Classroom Practice Questions

Environmental Engineering

Solutions for Volume : I Classroom Practice Questions

Water Supply Engineering

Chapter- 1
Population Forecasting & Water Demands

Class Room Practice Solutions

01. Ans: (c)

Sol:

Year	Population	Per decade increased in Population
1970	40000	
		6000
1980	46000	
		7000
1990	53000	
		5000
2000	58000	

$$\bar{x} = \frac{6000 + 7000 + 5000}{3} = 6000$$

$P_{2010} = P_0 + n\bar{x}$

$$n = \frac{2010 - 2000}{10}$$

$n = 1$

$P_{2010} = 58000 + 1 \times 6000$

$P_{2010} = 64000$

02. Ans: (c)

Sol:

Time	Population	Per decade % age increased in Population
1	1	
2	1.4	$\frac{1.4-1}{1}\times 100 = 40$
3	1.68	$\frac{1.68-1.4}{1.4}\times 100 = 20$

$\bar{r} = (r_1 \times r_2)^{1/2}$

$\bar{r} = (40 \times 20)^{1/2}$

$\bar{r} = 28.28\,\%$

$P_0 = 1.68\,\text{lakh}$

$$P_1 = P_0\left(1 + \frac{\bar{r}}{100}\right)^1$$

$$P_1 = 1.68\left(1 + \frac{28.28}{100}\right)^1$$

$P_1 = 2.15$ lakh $= 2.20$ lakh

03. Ans: 1.37, 1.97 billion

Sol: $K = 1.6\%$ per year

$P_{2000} = 1$ billion

$P_{2020} = ?$

$P_{2020} = P_{2000}.\ e^{k(2020\text{-}2000)}$

$$P_{2020} = 1 \times e^{\frac{1.6}{100}(20)}$$

P_{2020} = 1.37 billion

K = 1.2% per year

P_{2020} = 1.37 billion

P_{2050} = ?

$P_{2050} = P_{2020}.\ e^{k(2050-2020)}$

$P_{2050} = 1.37 \times e^{0.012\times30}$

P_{2050} = 1.97 billion

04. Ans: 743900

Sol: $\bar{x} = 5000$ per decade

$\bar{y} = 500$ per decade

P_{2020} = ?

P_{1990} = 50000 (given)

$$n = \frac{2020-1990}{10} = 3$$

$$P_{2020} = P_{1990} + 3\bar{x} + \frac{3(3+1)}{2}.\bar{y}$$

$$P_{2020} = 50000 + 3 \times 5000 + \frac{3\times4}{2}\times500$$

P_{2020} = 68000

05. Ans: 743900

Sol:

Year	Population	Per decade increased in Population	Incremental over increase in population
1960	250000		
		230500	
1970	480500		-160700
		69800	
1980	550300		+18500
		88300	
1990	638600		-31700
		56600	
2000	695200		

$$\bar{x} = \frac{230500+69800+88300+56600}{4}$$

$\bar{x} = 111300$

$$\bar{y} = \frac{-160700+18500-31700}{3}$$

$\bar{y} = -57966.67$

$$P_n = P_0 + n\bar{x} + \frac{n(n+1)}{2}.\bar{y}$$

P_{2020} = ?

$$P_0 = 695200, \quad n = \frac{2000-1980}{10} = 2$$

$$P_{2020} = 695200 + 2\times111300 + \frac{2(2+1)}{2}(-57966.67)$$

P_{2020} = 743900

Per capita water supply = 200 *l*pcd

$Q_{2020} = P_{2020}$ × per capita water supply

= 743900 × 200 × 10^{-6} MLD

= 148.7 MLD = 149 MLD

06. Ans: 1540000

Sol:

Time	Population	Per decade increased in Population	Per decade % age increased in population
1	400000		
		158500 (dP_1)	
2	558500		$\frac{158500}{400000}\times100 = 39.6\%$
		217500 (dP_2)	
3	776000		$\frac{217500}{558500}\times100 = 38.94\%$
		322500 (dP_3)	
4	1098500		$\frac{322500}{776000}\times100 = 41.5\%$
5	?		

If $dP_1 < dP_2 < dP_3 \rightarrow$ Geometric increase used

$\bar{r} = (39.6 \times 38.94 \times 41.5)^{1/3} = 39.99\%$

$= 40\%$

$$P_n = P_0\left(1+\frac{\bar{r}}{100}\right)^n$$

$$P_5 = 1098500\left(1+\frac{40}{100}\right)^1 \approx 1540000$$

07. Ans: (c)

Sol: $P_0 = 28,000$

Average increase per decade,

$$\bar{x} = \frac{44,000-28,000}{2} = 8,000$$

$4200 m^3/d$ required for 28,000 persons

$6000 m^3/d$ sufficient for ------- persons

$$= \frac{28000\times 6000}{4200} = 40,000$$

$P_n = P_0 + n\bar{x}$

$40,000 = 28,000 + n\times 8000$

$n = 1.5$ decades $= 15$ years

08. Ans: 100765

Sol:

Year	Population	Per decade Percentage Increase in Population	Decrease in Percentage Increase
1960	55,500	14.77	2.84
1970	63,700	11.93	0.43
1980	71,300	11.5	–
1990	79,500	–	–

Per decade percentage increase in population

$$= \frac{63700-55500}{55500}\times 100 = 14.77\%$$

$r_0 = 11.5$

$$P_n = P_o\left[1+\frac{r_o-\bar{D}}{100}\right]\left[1+\frac{r_o-2\bar{D}}{100}\right].....\left[1+\frac{r_o-n\bar{D}}{100}\right]$$

P_o = latest known population

P_n = prospective population after n year

r_o = latest per decade percentage increase in population

$\bar{D}$ = average decrease in percentage increase

$$\bar{D} = \frac{2.84+0.43}{2}$$

$\bar{D} = 1.635$

$$n = \frac{2020-1990}{10} = 3$$

$$P_n = P_o\left[1+\frac{r_o-\bar{D}}{100}\right]\left[1+\frac{r_o-2\bar{D}}{100}\right]\left[1+\frac{r_o-3\bar{D}}{100}\right]$$

$$P_n = 79500\left[1+\frac{11.5-1.635}{100}\right]\left[1+\frac{11.5-2(1.635)}{100}\right]\left[1+\frac{11.5-3(1.635)}{100}\right]$$

$= 79500\ (1.09865)\ (1.0823)\ (1.06595)$

$= 100765.29$

$\simeq 100765$

$\therefore$ Population for the year 2020 by decreasing rate of growth = 100765

Chapter- 3
Quality of Water

Class Room Practice Solutions

01. Ans: (d)

Sol: $Ca^{2+} = 160\,mg/\ell$

$Mg^{2+} = 40\,mg/\ell$

$$TH = Ca^{2+} \times \frac{50}{20} + Mg^{2+} \times \frac{50}{12}$$

$$TH = 160 \times \frac{50}{20} + 40 \times \frac{50}{12}$$

TH = 567 mg/*l* *as* $CaCO_3$

02. Ans: (d) **03. Ans: (d)**

04. Ans: (c)

Sol: $$FTN = \frac{A+B}{A} = \frac{25+175}{25}$$

FTN = 8

05. Ans: (d)

Sol: The product of H^+ ions and OH^- ions in a stronger acids = 10^{-14}

06. Ans: (d) **07. Ans: (d)**

08. Ans: (c) **09. Ans: (b)**

10. Ans: (b)

11. Ans: (a)

Sol: TH = 200 mg/*l* as $CaCO_3$

TA = 250 mg/*l* as $CaCO_3$

TH < TA

CH = TH = 200 mg/*l*

12. Ans: (d) **13. Ans: (a)**

14. Ans: (d)

Sol: $(PH)_I = 7.2$, $(H^+)_I = 10^{-7.2}$ mol/lit

$(PH)_0 = 8.4$, $(H^+)_0 = 10^{-8.4}$ mol/lit

$$\overline{H}^+ = \frac{(H^+)_I + (H^+)_0}{2}$$

$$\overline{H}^+ = \frac{10^{-7.2} + 10^{-8.4}}{2} = 3.3 \times 10^{-8} \text{ mol/lit}$$

$$P\overline{H} = \log_{10}\frac{1}{H^+} = \log_{10}\frac{1}{3.3\times10^{-8}}$$

$P\overline{H} = 7.47$

15. Ans: (c)

Sol: $(pH)_A = 4.4, (H^+)_A = 10^{-4.4}$ mol/lit

$(pH)_B = 6.4, (H^+)_B = 10^{-6.4}$ mol/lit

$$\frac{(H^+)_A}{(H^+)_B} = \frac{10^{-4.4}}{10^{-6.4}} = 100$$

16. Ans: (a)

Sol: Sample A, V_A = 300 m*l*

$(pH)_A = 7$

$(H^+)_A = 10^{-7}$ mol/lit

Sample B,

V_B = 700 m*l*

$(pH)_B = 5$

$(H^+)_B = 10^{-5}$ mol/lit

$$C_{mix} = \frac{V_A C_A + V_B C_B + ...}{V_A + V_B + ...}$$

$$(H^+)_{mix} = \frac{300\times10^{-7} + 700\times10^{-5}}{300+700}$$

$$(H^+)_{mix} = 7.03\times10^{-6} \text{ mol/lit}$$

$$(pH)_{mix} = \log_{10}\frac{1}{(H^+)_{mix}} = \log_{10}\frac{1}{7.03\times10^{-6}}$$

$$(pH)_{mix} = 5.15$$

17. Ans: (d)

Sol: $CO_3^{--} = 90$ mg/l

$HCO_3^- = 61$ mg/l

$$TA = CO_3^{--}\times\frac{50}{30} + HCO_3^-\times\frac{50}{61}$$

$$TA = 90\times\frac{50}{30} + 61\times\frac{50}{51}$$

TA = 200 mg/l as $CaCO_3$

18. Ans: (d)

Sol: From 10 –1–0.10 (MPN) against 4 –3–1, +Ve grouping

MPN = 33

For 1 – 0.1–0.01 dilution against 4 –3 –1 +ve group

$$MPN = 33\times\frac{\text{Table dilution}}{\text{Test dilution}}$$

MPN = 33 × 10 = 330 no/100 ml

19. Ans: (a)

Sol: PH = 9

$H^+ = 10^{-9}$ mol/lit

$OH^- = 10^{-5}$ mol/lit

OH^- mol/lit

$$OH^- \text{ mol/lit} = OH^- \times \text{Mol. wt. of } OH^- \times 1000$$

$$OH^- = 10^{-5}\times17\times1000 = 0.17 \text{ mg/lit}$$

$$OH^- = 0.17\times\frac{50}{17} = 0.50 \text{ mg/lit as } CaCO_3$$

20. Ans: (d) **21. Ans: (d)**

22. Ans: (a) **23. Ans: (b)**

24. Ans: (b) **25. Ans: (c)**

26. Ans: (d)

27. Ans: (d)

Sol: TA = 250 mg/l as $CaCO_3$

TH= 350 mg/l as $CaCO_3$

TH > TA

CH = TA = 250 mg/l

NCH = TH – CH = 350 –250

NCH = 100 mg/l as $CaCO_3$

28. Ans: (c)

Sol: $$TH = Ca\times\frac{50}{20} + Mg\times\frac{50}{12.2}$$

$$TH = 55\times\frac{50}{20} + 10\times\frac{50}{12.2} = 178.48$$

TH ≃ 179 mg/l as $CaCO_3$

Common data for Q 29 & 30

29. **Ans: (a)**

Sol: TH in mg/lit as C_aCO_3

$$= Ca^{++} \text{ in mg/lit} \times \frac{50}{20} + Mg^{++} \text{ in mg/lit} \times \frac{50}{12}$$

$$= (12 \times 20) \times \frac{50}{20} + (18 \times 12) \times \frac{50}{12}$$

= 1500 mg/lit as $CaCO_3$

30. **Ans: (c)**

Sol: Alkalinity in mg/lit $CaCO_3$

$$= HCO_3^- \text{ in mg/lit} \times \frac{50}{61} + CO_3 \text{ in mg/lit} \times \frac{50}{30}$$

$$= (30 \times 61) \times \frac{50}{61} + (5 \times 30) \times \frac{50}{30}$$

= 1750 mg/lit as $CaCO_3$

Common data for Question Nos. 31. & 32

31. **Ans: (c)**

Sol: $TH = 100 \times \frac{50}{20} + 6 \times \frac{50}{12}$

= 275 mg/lit as $CaCo_3$

32. **Ans: (a)**

Sol: $\text{Alkalinity} = 250 \times \frac{50}{61}$

$= 204.9 \approx 205 \text{mg/lit } as\ CaCO_3$

33. **Ans: (d)**

Sol: Tomoto juice pH = 4.1

$$pH = \log_{10} \frac{1}{H^+}$$

$$4.1 = \log_{10} \frac{1}{H^+}$$

$H^+ = 10^{-4.1}$ mol/lit

$H^+ = 7.94 \times 10^{-5}$ mol/lit

34. **Ans: (d)**

Sol: $OH^- = 10^{-5.6}$ m.mol/lit

$OH^- = 10^{-5.6} \times 10^{-3}$ mol/lit

$OH^- = 10^{-8.6}$ mol/lit

$$H^+ = \frac{10^{-14}}{10^{-8.6}} = 10^{-5.4}$$

$$pH = \log_{10}\left(\frac{1}{H^+}\right) = \log_{10} \frac{1}{10^{-5.4}}$$

pH = 5.4

35. **Ans: (b)**

36. **Ans: (d)**

Sol: From table 10–1–0.10 (MPN) against

2–1–0 + ve group

MPN = 7

For 1–0.1 –0.01 dilution against

2–1–0, +ve group

$$MPN = 7 \times \frac{\text{table dilution}}{\text{test dilution}}$$

MPN = 7 ×10 = 70

37. Ans: (b)

Sol: Mol. Wt of $CO_3 = 12 + 3\times16 = 60$

Mol. Wt. of Ca = 40

60 parts of CO_3 required = 40 parts of Ca

1 part of CO_3 require $= \dfrac{40}{60}$ part of Ca

90 mg/l part of CO_3 require $= \dfrac{40}{60}\times 90$ mg/l of Ca

= 60 mg/l of Ca

38. Ans: (c)

Sol: PH = 9.25

$$PH = \log_{10}\frac{1}{H^+}$$

$$9.25 = \log_{10}\frac{1}{H^+}$$

$$H^+ = 10^{-9.25}\ \text{mol/lit}$$

$$[H^+][OH^-] = 10^{-14}$$

$$[OH^-] = \frac{10^{-14}}{10^{-9.25}} = 10^{-4.75}\ \text{mol/lit}$$

$$OH^-(mg/\ell) = OH^-(mg/1) \times \text{Mol. Wt . of } OH^- \times 1000$$

$$OH^-_{(mg/\ell)} = 10^{-4.75}\times 17\times 1000$$

OH^- = 0.302 mg/l as $CaCO_3$

39. Ans: (d)

Sol: $TON = \dfrac{A+B}{A}$

$$TON = \frac{187.5+12.5}{12.5}$$

TON = 16

40. Ans: (b)

Sol: OH^- = 17 mg/l

$$OH^-(mol/l) = \frac{(OH^-)mg/1}{\text{Mol. wt. of } OH^- \times 1000}$$

$$OH^- = \frac{17}{17\times1000} = 10^{-3}\ \text{mol/lit}$$

$$[H^+][OH^-] = 10^{-14}\ \text{mol/lit}$$

$$H^+ = \frac{10^{-14}}{10^{-3}} = 10^{-11}\ \text{mol/lit}$$

$$pH = \log_{10}\frac{1}{H^+} = \log_{10}\frac{1}{10^{-11}}$$

pH = 11

41. Ans: (b)

Sol: Ca^{2+} = 4 m.eq/lit

Mg^{2+} = 1 m eq/lit

HCO_3^- = 3.5 m. eq/lit

TH = Ca^{2+} * 50 + Mg^{2+} * 50

(where Ca & Mg are in m.eq/lit)

= 4×50+1×50 = 250 mg/l as $CaCO_3$

TA = CO_3 * 50 + HCO_3 *50

(where CO_3 & HCO_3 in m.eq/lit)

= 0+3.5×50 = 175 mg/l as $CaCO_3$

TH > TA

∴ CH = TA = 175 mg/l as $CaCO_3$

NCH = TH – CH = 250 – 175

= 75 mg/l as $CaCO_3$

42. Ans: (b) **43. Ans: (d)**

44. Ans: 640, 220 & 420 mg/lit

Sol: W_1 = 98.42 gm

W_2 = 98.484 gm

W_3 = 98.462 gm

(i) Total solids $(TS) = \dfrac{W_2 - W_1}{V}$

$$TS = \frac{98.484 - 98.42}{100} \times 10^6$$

$= 640$ *mg/l*

(ii) Volatile solids $= \dfrac{W_2 - W_3}{V}$

$$= \frac{98.484 - 98.462}{100} \times 10^6$$

= 220 mg/*l*

(iii) Fixed solids $= \dfrac{W_3 - W_1}{V}$

$$= \frac{98.462 - 98.42}{100} \times 10^6$$

= 420 mg/*l*

Chapter- 4
Plain Sedimentation

Common data for Qs. 1& 2

01. Ans: (b)

Sol: B = 6 m, L = 15 m, H = 3 m

Q = 2 MLD

Surface over flow rate $V_0 = \dfrac{2 \times 10^6}{6 \times 15 \times 24}$

$= 926$ lit/hr/m^2

02. Ans: (d)

Sol: Detention time is

Volume of setting tank = Q × D.T

$$DT = \frac{6 \times 15 \times 3}{\dfrac{2 \times 10^6}{24 \times 10^3}} = \frac{270}{83.33} = 3.24 \text{ hrs}$$

Common data question for Q3 and Q4

03. Ans: (c)

Sol: **Q =** 1.8 MLD

D.T = 4 hours

V = ?

$$= \frac{1.8 \times 10^6}{10^3 \times 24} = 75 \text{ m}^3/\text{hr}$$

Volume of tank = Q×DT

$= 75 \times 4 = 300$ m^3

04. Ans: (b)

Sol: SOR = 500 lit/hr/m, L:B = 4:1, L =?

$$V_0 = \frac{Q}{A}$$

$$\text{Surface area} = \frac{Q}{SOR} = \frac{1.8 \times 10^6}{24 \times 500}$$

L × B = 150 m^2

$$L \times \frac{L}{4} = 150 \text{ m}^2$$

$L^2 = 150 \times 4$

L = 24.49 m

05. Ans: 0.0112 m/sec

Sol: V_H =? L = 60 m, H = 3 m

$$\frac{L}{V_H} = \frac{H}{V_S}$$

$$V_S = \frac{g(s-1)d^2}{18\nu}$$

$$= \frac{9.81(2.65-1)\times(0.025\times10^{-3})^2}{18\times\frac{0.01}{(100)^2}}$$

$= 5.62 \times 10^{-4}$ m/sec

$$\frac{60}{V_H} = \frac{3}{5.62\times10^{-4}}$$

$\Rightarrow V_H = 0.0112$ m/sec

06. Ans: (c)

Sol: Q = 100000 m^3/day

Settling velocity = 20 m /day

$$\text{Area of tank } = \frac{Q}{\text{Settling velocity}}$$

$$\text{Area of tank } = \frac{100000}{20} = 5000\,m^2$$

07. Ans: (a)

Sol: d = 0.025 mm, s = 2.65, ν=0.01 cm^2/sec

$$\nu_s = \frac{g(s-1)d^2}{18\times\nu}$$

$$= \frac{9.81(2.65-1)\times(0.025\times10^{-3})^2}{18\times0.01\times10^{-4}}$$

= 0.056 cm/sec

08. Ans: 27.08, 100%

Sol: V_0 = 12,000 lit/hr/m^2, d = 0.03 mm

= 12 m^3/hr/m^2

$$V_S = \frac{g(s-1)d^2}{18\times\nu}$$

$$= \frac{9.81(2.65-1)(0.03\times10^{-3})^2}{18\times\frac{0.897}{(1000)^2}}$$

V_S = 0.922 mm/sec

$V_S = 9.02\times10^{-4}$ m/sec

$$\eta = \frac{V_S}{V_0}\times100$$

V_0 = 12 m^3/hr/m^2 = 3.33×10^{-3} m/sec

$$= \frac{9.02\times10^{-4}}{3.33\times10^{-3}}\times100$$

= 27.08%

b) d = 12 mm > 1 mm

$$V_s = 18\sqrt{g(G-1)d}$$

$$= 18\sqrt{9.81(2.65-1)\times((12\times10^{-3}))}$$

$$\eta = \frac{V_S}{V_0}\times100 = \frac{0.79\times100}{3.33\times10^{-3}}\times100$$

= 100%

Common Data for Qs. 09 & 10

09. Ans: (a)

Sol: d = 26 m with H = 2.10 m

Q = 13000 m^3/day, D.T = ?

$$DT = \frac{2.10\times\left(\frac{\pi}{4}\times26^2\right)}{\frac{13000}{24}} = 2.05\text{ hrs}$$

10. Ans: (d)

Sol: Weir loading $= \dfrac{Q}{\text{length of weir}}$

$$= \frac{13000}{\pi \times 26}$$

$$= 159 \text{ m}^3/\text{day/m}$$

11. Ans: 44.5, 14.83 & 0.47

Sol: $d = 50\ \mu m = 50 \times 10^3$ m = 0.05 m,

G = 2.3, Q = 100 MLD, H = 3m

L:B = 3:1, d = 3 m, $\nu = 1.01 \times 10^{-6}$ m^2/s

$$V_S = \frac{g(G-1)d^2}{18 \times \nu}$$

$$= \frac{9.81(2.3-1)\times(0.05\times10^{-3})^2}{18\times1.01\times10^{-6}}$$

$V_S = 1.753 \times 10^{-3}$ m/sec

For 100% removal

$V_S = V_0 = 1.753 \times 10^{-3}$ m/sec

$$\text{Surface area} = \frac{Q}{V_0} = \frac{1.157}{1.753\times10^{-3}}$$

L : 3B = 660.011

L × B

3B × B = 660.011

$3B^2 = 660.011$

$$B = \sqrt{\frac{660.011}{3}} = 14.83 \text{ m}$$

L = 3 × B = 3 × 14.83 = 44.49 m

$$D.T = \frac{\text{Volume of tank}}{Q} = \frac{L\times B\times H}{Q}$$

$$\frac{44.49\times14.83\times3}{\frac{100\times10^6}{24\times10^3}} = 0.47 \text{ hr}$$

12. Ans: 20 m^3/m^2/day

Sol: L × B × h = 100 × 50 × 3 m,

Q = 1,00,000 m^3/day

Surface flow rate -- ?, $\rho = 2.65$ g/cc,

$\nu = 1.02 \times 10^{-2}$cm^2/sec

$$\text{Surface flow rate} = \frac{Q}{\text{surface flow area}}$$

$$= \frac{100000}{100\times50} = 20 \text{ m}^3/\text{m}^2/\text{day}$$

For 100% removal $V_S = V_0$

= 20 × m^3/m^2/day

$$= \frac{20}{24\times60\times60} = 2.31 \times 10^{-4}$$

$$S = \frac{\rho_P}{\rho_w} = \frac{2.65}{1} = 2.65$$

$$V_S = \frac{g(s-1)d^2}{18\times v}$$

$$2.31 \times 10^{-4} = \frac{9.81(2.65-1)\times d^2}{1.836\times10^{-5}}$$

$d = 1.61 \times 10^{-5}$ mm

$d = 1.61 \times 10^{-2}$ m

13. Ans: (c)

Sol: H = 3 m, surface area = 900 m^2,

Q = 8000 m^3/day, T = 20^{0}C,

$\mu = 10^{-3}$ kg/m-s, $\rho = 1000$ kg/m^3,

d = 0.01 mm, G = 2.65, η = ?

$$V_o = \frac{Q}{\text{S.area}} = \frac{8000}{900} = 8.889 \text{ m}^3/\text{day/m}^2$$

$$= \frac{8.8}{24\times60\times60} \text{ m/sec} = 1.0185\times10^{-4}$$

$$V_s = \frac{g[\rho_p - \rho_w]d^2}{18}$$

$$= \frac{9.81(2650-1000)\times(0.01\times10^{-3})^2}{18\times10^{-3}}$$

$V_0 = 8.99 \times 10^{-5}$ m/sec

Proportion of particle removed 'η'

$$= \frac{V_S}{V_0}\times100$$

$$= \frac{8.99\times10^{-5}}{1.018\times10^{-4}}\times 100 = 88.31\%$$

Common data for Question Nos. 14 & 15

14. Ans: (a)

Sol:

L = 20 m, B = 10 m, H = 3 m, Q = 4 MLD,

T = 20^0C, $\mu = 1.002 \times 10^{-3}\ \frac{N-s}{m^2}$ at 20^0C,

$\rho_w = 998.2$ kg/m^3, G = 2.65

$$\text{Surface overflow} = \frac{Q}{\text{Surface area}}$$

$$= \frac{4\times10^6\times10^3}{20\times10}$$

$= 20$ m^3/m^2/day

15. Ans: (b)

Sol: η = 100%

$$V_s = \frac{g(\rho_p - \rho_w)d^2}{18\mu}$$

G = 2.65, $\rho_p = 2.65 \times \rho_w$

$= 2.65 \times 998.2$

$\rho_p = 2645.23$ kg/m^3

$$\frac{20}{24\times60\times60} = \frac{9.81(2645.23-998.2)d^2}{18\times1.002\times10^{-3}}$$

d = 0.016 mm

16. Ans: (b)

Sol: $V_{S1} = 0.1$ mm/s, $V_{S2} = 0.2$ mm/s,

$V_{S3} = 1.0$ mm/s

Surface over flow rate = 43.2 m^3/m^2/d

Particle	Percentage	V_S mm/sec	$\eta = \frac{V_S}{V_0}\times100$
1	10	0.1	$\eta_1 = \frac{0.1}{0.5}\times100 = 20$
2	60	0.2	$\eta_2 = \frac{0.2}{0.5}\times100 = 40$
3	30	1.0	$\eta_3 = \frac{1}{0.5}\times100 = 20$

Overal removal = ΣPini

$$= \frac{10}{100}\times20 + \frac{60}{100}\times40 + \frac{30}{100}\times100$$

= 56%

17. Ans: (b)

Sol: $V_0 = 30$ m^3/m^2/day, S = 2.65

$\rho = 1000$ kg/m^3

$\mu = 0.001$ N-s/m^2, 1 stoke = 10^{-4}m^2/sec

$\rho_p = s \times \rho_w = 2.65 \times 1000 = 2650$

$\rho_w = 1000$ kg/m^3

$$V_S = \frac{g(\rho_p - \rho_w)d^2}{18\mu}$$

$$= \frac{9.81(2650-1000)d^2}{18\times 0.001}$$

$$= \frac{30\times 10^3}{24\times 60\times 60} = \frac{16186.5d^2}{0.018}$$

$0.540 = 1.39851 \times 10^9 d^2$

$d = 1.965 \times 10^{-5}$

$d = 0.02$ mm

18. Ans: 3.1214

Sol: $d = 0.06$ mm $= 0.06 \times 10^{-3}$ m

$g = 9.8$ m/sec^2

$G = 2.65$

$\upsilon = 1.0105\times 10^{-2}$ cm^2/sec

$$V_s = \frac{g(G-1)d^2}{18\upsilon}$$

$$V_s = \frac{9.81\times(2.65-1)\times(0.06\times 10^{-3})^2}{18\times 1.0105\times 10^{-6}}$$

$V_s = 3.20\times 10^{-3}$ m/sec

$$\text{Surface area} = \frac{Q}{V_s}$$

$$A = \frac{0.01}{3.20\times 10^{-3}} = 3.124\,\text{m}^2$$

19. Ans: 22.576

Sol: $\eta = \frac{V_s}{V_o}\times 100$

$V_o = 40$ m^3/m^2/day $= 40$ m/day

$= 4.629 \times 10^{-4}$ m/sec

$$\eta = \frac{V_s}{V_o}\times 100$$

$$\Rightarrow 90 = \frac{V_s}{4.629\times 10^{-4}}\times 100$$

$V_s = 4.166 \times 10^{-4}$ m/sec

$$V_s = \frac{g[s-1]d^2}{18\gamma}$$

$$\Rightarrow 4.166\times 10^{-4} = \frac{9.81[2.65-1]}{18\times 1.1\times 10^{-6}}d^2$$

$d = 22.576 \times 10^{-6}$ m

$= 22.576$ μm

20. Ans: 112.66

Sol: Surface over flow rate

$V_o = 32.5$ m^3/day/m^2

$L = 32.5$ m, $B = 8.0$ m, $D = 2.25$ m

$$L\times B = \frac{Q}{V_o}$$

$$\Rightarrow 32.5\times 8 = \frac{Q}{32.5}$$

$Q = 32.5 \times 8 \times 32.5$ m^3/day

$$\text{Weir loading rate} = \frac{Q}{\text{Length of weir}}$$

$$= \frac{32.5\times 8\times 32.5}{75}$$

$= 112.66$ m^3/day/m

Chapter- 5
Coagulation

01. Ans: (d)

Sol: Q = 28800 m^3/d; ρ_w = 1000 kg/m^3

$\upsilon = 10^{-6}$ m^2/sec; G = 900 s^{-1}

DT = 2 min

Volume of mixing basin = Q × DT

$$V = \frac{28800\times 2}{24\times 60} = 40\,m^3$$

$$G = \sqrt{\frac{P}{V\mu}}$$

$$900 = \sqrt{\frac{P}{40\times 10^{-6}\times 1000}}$$

P = 32400 watts

Common Data Question 2 & 3

02. **Ans: (c)**

Sol: Q = 10MLD = 10× 10^6 Lit/day

Alum = 20 mg/Lit

1 mg of Alum requires 0.45 mg of Alkaline as $caco_3$

∴ 20 mg/Lit of alum requires

= 20×0.45 = 9 mg of alkaline as $caco_3$ per Lt of water

∴Total alkalinity matching filter

Alum = 9 mg/Lit

= 10×10^6 Lit /day = 90×10^6

Total alkanity requirement (10^6 mg per day)

= 90

03. **Ans: (d)**

Sol: Natural available alkalinity = 6 mg/ Lit

∴ Alkalinity to be added additionally

= 9 – 6 = 3 mg/Lit

∴ Alkaline to be added to the water

= 3 × 0.56 = 1.68 mg/Lt

Total quick lime required per year

$$= \frac{1.68\times 10\times 10^6\times 365}{10^6}$$

Total quick lime required (10^6 mg per year)

= 6132

04. Ans: 168 kg/day, 5.55 mg/*l*

Sol: Alum required in order to total = 12 MLD

Alum dose requirement = 14 ppm

Co_2 gas = ?

Total alum requirement/day

= Q × alum dose

= 12 × 14 = 168 kg/day

$Al_2(So_4)_3 = 18\ H_2o + Ca(HCO_3)_2$

$\rightarrow 3CaSo_4 + 2Al(OH)_3 + 6CO_2 + 18H_20$

Molecular weight of alum = 666

Molecular weight of $6CO_2 = 6[c(O_2)]$

= 6[12 + 2 × 16]

= 264

666 parts alum release = 264 parts of CO_2

1 part alum release = $\frac{264}{666}$ parts of CO_2

14 mg/l of alum release = $\frac{264}{666}$ × 14 mg/*l* of CO_2

= 5.54 mg/*l* of CO_2

05. Ans:120 kg/day, 24.168 kg/day

Sol: Q = 12 MLD

Dosage of ferrous sulphate 10 mg/*l*

Total quantity of ferrous sulphate and lime-?

$= FeSO_4 . 7H_2O$

Total quantity of ferrous sulphate req/day

= Q × dosage of ferrous sulphate

= 10 × 12

= 120 kg/day

$FeSO_4\ 7H_2O + Ca(OH)_2$

$\rightarrow CaSO_4 + Fe(OH)_2 + 7H_2O$

$Ca(OH)_2 \rightarrow CaO + H_2O$

Molecular weight of $FeSO_4$.

$7H_2O$ = Fe + S + 4(O) + 7(2+H+O]

Molecular weight of CaO = 56

278 parts of ferrous sulphate required

= 56 parts of CaO

1 part of ferrous sulphate required

$= \frac{56}{278}$ parts of CaO

10 mg/l of ferrous $= \frac{56}{278} \times 10$ mg/*l* as CaO

= 2.14 mg/*l*

Total lime as CaO required/day

= Q × dose of CaO

= 12 × 2.014

= 24.168 kg/day

06. Ans:1613.92 watts

Sol:

Q = 3000 m^3/hr; G = 40 sec^{-1}

D.T = 20 min; $\mu = 1.0087 \times 10^{-3}$ N.s/m^2

Volume = Q × DT = $3000 \times \frac{20}{60} = 1000$ m^3

$\frac{L}{B} = 2$ D = 0.40B

L = 2B

Surface area $= \frac{\text{Volume}}{\text{depth}} = \frac{1000}{0.40B}$

$L \times B = \frac{1000}{0.40B}$

B × 2B × 0.40B = 1000

$0.80B^3 = 1000$

B = 10.77 m, L = 21.54 m, D = 4.3 m

$G = \sqrt{\frac{P}{v.\mu}}$

$40 = \sqrt{\frac{P}{1000 \times 1.008 \times 10^3}}$

P = 1613.92 watts

07. Ans: (a)

08. Ans: (b)

Sol: $G = \sqrt{\frac{p}{V\mu}}$

$600 = \sqrt{\frac{p}{2 \times 1 \times 10^{-3}}}$

$P = (600)^2 \times 2 \times 1 \times 10^{-3}$ = 720 watts

09. Ans: (b)

Sol:

$Q = 4.2\ m^3/min$

$V_0 = 0.2$ mm/sec;

$d = 3.5$ m

$$V_0 = \frac{Q}{\text{surface area}}$$

$$0.2 \times 10^{-3} = \frac{4.2}{60 \times A}$$

$A = 350\ m^2$

10. Ans: 37800 kg

Sol: $Q = 35\ m^3/min = 50.4$ MLD

Dose of alum = 25 mg/l

Total alum required in kg/day

= Q (MLD) × dose of alum (mg/l)

= 50.4 ×25 = 1260 kg/day

Monthly alum requirement = 1260 × 30

= 37800 kg

Chapter- 6
Filtration

01. Ans: 35.35 m, 17.67 m

Sol: P = 50,000 person

ROF = 180 lit/hr/m^2

Q = Population × per capita demand

Q = 50,000 × 150 = 7500000 lit/day

Design discharge $Q_{density} = 1.8 \times Q$

= 1.8 × 7500000

= 135×10^5 lit/day

ROF = 180 lit/hr/m^2

$$\text{Total area of slow sand filter} = \frac{Q_{density}}{ROF}$$

$$= \frac{135\times10^5}{24\times180}\ \frac{\text{lit/hr}}{\text{lit/hr/m}^2} = 3125\ m^2$$

No. of filters in operations = 5

(1 act as stand by)

Area of each filter L × B

$$= \frac{\text{Total area}}{\text{no. of filters in operation}}$$

$$= \frac{3125}{5} = 625 m^2$$

$L : B = 2 : 1 \Rightarrow L = 2B$

$2B \times B = 625\ m^2$

$B^2 = 625/2$

$$B = \sqrt{\frac{625}{2}} = 17.67 \text{ m}$$

L = 35.35 m

02. Ans: 11

Sol: Treated water Q = 10 MLD

$= 10 \times 10^3\ m^3/day$

Rate of filtration RDF = 200 lit/hr/m²

$= 200 \times 10^{-3} \times 24\ m^3/day/m^2$

Total area required to treat water

$$= \frac{Q}{RDF} = \frac{10\times10^3}{200\times10^{-3}\times24}$$

= 2083.33 m^2

Area of each filter = 20 ×10 M^2

$$\text{No. of filters required} = \frac{\text{Total area of filters}}{\text{area of each filter}}$$

$$= \frac{2083.33}{200}$$

$$= 10.41 \simeq 11 \text{ numbers}$$

03. Ans: 20 m, 10 m, 4.19%

Sol: Q = 24 MLD, ROF = 5 $m^3/hr/m^2$,

L.B = 2 : 1

Total area of RSF req $= \frac{1000}{5} = 200m^2$

$$Q = \frac{24\times10^6\times10^{-3}}{24} = 1000m^3/hr$$

No. of filters = 1

Area of filter = 200 m^2

L: B = 200

2B × B = 200

$$B^2 = \frac{200}{2} = \sqrt{\frac{200}{2}} = 10 \text{ m}$$

L = 2 × 10 = 20 m

Volume of water filter b/w back wash = ROF × duration of filtration × area of each filter

$$= 5\times\left(24-\frac{10}{60}\right)\times200 = 23833.33m^3$$

Volume of water used in back wash ROB × DOB × area

$$= 6\times5\times\frac{10}{60}\times200 = 1000 \text{ m}^3$$

% of filter water used in back wash

$$= \frac{\text{volume of water filtered back wash}}{\text{volume of filtered b/w back}}$$

$$= \frac{1000}{23833.3}\times100 = 4.19\%$$

04. Ans: 9.48 m, 4.74 m, 0.225 m^3/sec

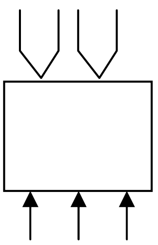
Sol: Q = 0.25 m^3/sec, No. of filters = 4,

ROF = 5m^3/m^2/hr

Back wash rater, 10 l/m^2/sec,

L : B = 2 : 1

L = ?, B = ?



$$\text{Total area of RSF} = \frac{Q}{ROF} = \frac{0.25}{\frac{5}{60\times60}}$$

$$= \frac{0.25\times60\times60}{6} = 180 \text{ m}^2$$

Area of each filter $= \frac{180}{4} = 45 \text{ m}^2$

L × B = 45

2B × B = 45

$2B^2 = 45$

$$B = \sqrt{\frac{45}{2}} = 4.74 \text{ m}$$

L = 2 × 4.74 = 9.48 m

= 10 × 45 = 450 liter/sec = 0.45 m^3/sec

Back wash water flow rate OB = velocity (V_B) × Area of each filter

= 10 × 45 = 450 lit/sec

= 0.45 m^3/sec

There are two troughs

Flow through each wash water $= \frac{Q}{2}$

$$= \frac{0.45}{2} = 0.225 \text{ m}^3/\text{sec}$$

05. Ans: 0.27 m

Sol: $V_1 = i_r = 30$ m/hr

$L = 0.6$ m $\quad \phi = 0.8$

$d = 0.5$ mm $\quad n = 0.4$

$S = 2.68 \quad \nu = 1\times10^{-6}$ m^2/sec

$$Re = \frac{(V_s.d)\phi}{\nu}$$

$$= \frac{3\times(0.5\times10^{-3})\times0.8}{1\times10^{-6}\times60\times60} = 0.33$$

$$f = \frac{150(1-n)}{Re} + 1.75$$

$$= \frac{150(1-0.4)}{0.33} + 1.75$$

$$= 271.75$$

$$h_f = \frac{f.L.V_s^2}{gd}\times\frac{(1-n)}{n^3\times\phi}$$

$$= \frac{271.75\times0.6\left(\frac{3}{60\times60}\right)^2}{9.81\times0.5\times10^{-3}}\times\frac{(1-0.4)}{0.4^3\times0.8}$$

$$= 0.27 \text{ m}$$

06. Ans: 0.032 m/sec, 0.6258 m

Sol: $d = 0.65$ mm, $G = 2.66$, $n = 0.42$,

$z = 65$ cm $= 0.65$ mm

$$V_s = \frac{g(s-1)d^2}{18\gamma}$$

$$= \frac{9.81(2.66-1)\times(0.65\times10^{-3})^2}{18\times1.3\times10^{-2}\times10^{-4}}$$

$$= 0.29 \text{ mm/sec}$$

$$V_B = V_S(n_e)^{4.5}$$

$$\frac{z_e}{z} = \frac{1-n}{1-n_e}$$

$$\frac{1.53}{z} = \frac{1-0.42}{1-n_e}$$

$1.5(1-n_e) -1-0.42 \Rightarrow 1.5 - 1.5\, n_e = 0.58$

$$n_e = \frac{1.5-0.58}{1.5}$$

$n_e = 0.613$

$V_B = 0.29(0.613)^{4.5} = 0.032$ m/sec

Head loss during back wash h_b

$= z(1-n)(s-1)$

$= 0.6258$ m

07. Ans: (a)

Sol: $$\frac{1}{z},\ \ln\frac{100}{100-\eta_1} = \frac{1}{z_2}\ln\frac{100}{100-\eta_2}$$

$Z_1 = 0.05$ m, $\eta_1 = 90\%$, $\eta_2 = 99\%$

$Z_2 = ?$

$$\frac{1}{0.05}\ln\frac{100}{100-90} = \frac{1}{z_2}\ln\frac{100}{100-99}$$

$$46.06 = \frac{1}{z_0}\times4.606$$

$$\Rightarrow z_2 = \frac{4.606}{46.06}$$

$Z_2 = 0.10$ m

Common data for Question Nos. 08 & 09

08. Ans: (c)

Sol: ROF = 200 $m^3/m^2/d$,

$Q = 0.5\ m^3/s$, $A = 50\ m^2$

$$\text{Total area} = \frac{Q}{ROF} = \frac{0.5}{\frac{200}{24\times60\times60}}$$

$= 216\ mm^2$

09. Ans: (c)

Sol: No. of filters = $\frac{216}{50} = 4.32 \backsimeq 5$

Total no. of filters = 6

10. Ans: 144 m^3/day/m^2

Sol: Q = $1m^3$/sec = 86400 m^3/day

no.of filters = 14

surface area of each filter = $50m^3$

no. of filters in working condition

= 14 – 2=12

$$\text{Loading rate} = \frac{Q}{\text{surface area}} = \frac{86400}{12\times50}$$

= 144 m^3/day/m^2

Chapter- 7
Disinfection

01. Ans: 51.2 sec

Sol: $N_o = 10^6$

$N_t = 100$

$$\ell n\left(\frac{100}{10^6}\right) = -6\times10^{-3}\times30\times t$$

t = 51.16 sec

02. Ans: (c)

Sol:

Population = 20,000 at a per capita demand

= 150 lit/day

Q = 20,000 × 150

$$= \frac{3000000}{10^6}\ \text{lit/day} = 3\ \text{MLD}$$

Bleaching powder = 0.2 mg/*l*

Cl_2 = 0.2 ppm = 0.2 mg/*l*

$Ca(Ocl)_2$ = 30% available Cl_2

Bleaching powder

$$= \frac{Cl_2\ \text{dose}}{\%\text{ of } Cl_2\ \text{bleaching powder}}$$

$$= \frac{0.2}{\frac{30}{100}} = 0.66\ \text{mg}/l$$

Total bleaching powder required/day

= 3 × 0.66 = 1.98 = 2 kg/day

03. Ans: (c)

Sol: Cl_2 = 20,000 Cu m

8 kg. residual after 10 min = 0.15 mg/l

Q = 20,000 m^3/day

Total Cl_2 used = 8 kg/day

Residual Cl_2 = 0.15 mg/*l*

Dose of Cl_2 = ?

Demand = ?

Total Cl_2 = Q × design of Cl_2

8 = 20,000 × dose of Cl_2 MLD

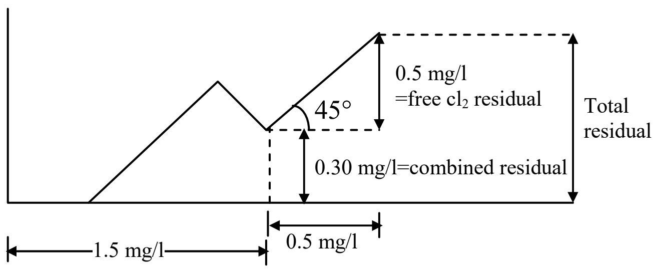
Dose of Cl_2 = $\frac{8}{20}$ = 0.4 mg/*l*

Cl_2 demand = Cl_2 dose – Residual Cl_2

= 0.4 – 0.15 = 0.25 mg/*l*

04. **Ans: (b)**

Sol:



Total residual = 0.3 + 0.5 = 0.80 mg/*l*

05. Ans: (b)

Sol: $$\frac{HOCL}{(HOCL+OCL)} \quad \frac{1}{1+\frac{K_i}{H^+}}$$

$$0.9 = \frac{1}{1+\frac{2.7\times10^{-8}}{H^+}}$$

$H^+ = 2.43 \times 10^{-7}$

$$P^H = \log_{10}\left[\frac{1}{H^+}\right]$$

$\Rightarrow P^H = 6.614$

06. Ans: (d)

07. Ans: (d) **08. Ans: (d)**

09. Ans: 3.2 min

Sol: Residual = 0.6 mg/l, K = 3×10^{-2} per sec

t = ?

$$\% \text{ of kill} = \frac{N_o - N_t}{N_o} \times 100$$

$99.7 -(1-e^{-kt})\times100$

$99.7 = \left(1 - e^{-3\times10^{-2}} \times t\right) \times 100$

$99.7 = (0.029 \times t) \times 100$

$$t = 193.63 \text{ sec} = \frac{193.63}{60}$$

= 3.2 min

10. Ans: (c)

Sol: Percent removal $(\%R)$

$$=\left(\frac{10^6-10^2}{10^6}\right)\times100=99.99\%$$

$\log(R)=\log10^6 - \log10^2 = 6\text{-}2 = 4$

11. Ans: (a)

Sol: Free residual

$= HOCL + OCL^- = 2mg/lit \text{ as } Cl_2$

$2(mg/lit) = Cl_2\left(\frac{moles}{lit}\right) \times Mol.wt \times 1000$

$\frac{2}{2\times 35.5\times 1000} = Cl_2 \ (moles/lit)$

$Cl_2 = 2mg/lit = 2.816\times 10^{-5}$ (moles/lit)

$P_H = 7.5 \quad \therefore P_H = \log_{10}\left(\frac{1}{H^+}\right)$

$\Rightarrow H^+ = 10^{-7.5}$ moles/lit

$\frac{HOCL}{OCL^-} = [K][H^+] = 10^{7.5}\times 10^{-7.5}$

$\therefore HOCL = OCL^-$

$\Rightarrow HOCL + OCL^- = 2mg/lit$

$= 2.816\times 10^{-5}$ (moles/lit)

$\because 2\,OCL^- = 2.816\times 10^{-5}$

$\because OCL^- = 1.408\times 10^{-5}$ moles/lit

12. Ans: (b)

Sol: Fraction HOCL = $\frac{1}{1+\frac{k}{H^+}}$

pH = 7 $\Rightarrow H^+ = 10^{-7}$ mole/lit

k = 2.5×10^{-8} mol/lit

$= \frac{1}{1+\frac{2.5\times 10^{-8}}{10^{-7}}} = \frac{1}{1.25} = 0.8$

13. Ans: (a)

Sol: $C^n t$ = constant $\Rightarrow t \propto \frac{1}{C}$

C $\rightarrow$ Concentration of disinfectant

t $\rightarrow$ detention time (or) contact time

n $\rightarrow$ dilution factor.

Contact time "t" = $\frac{\text{Volume of contact unit}}{\text{Flow rate}}$

$= \frac{V}{Q}$

$\therefore C^n t$ = constant

$\therefore C_1^n t_1 = C_2^n t_2$ = constant

$C_1^n \frac{V}{Q_1} = C_2^n \frac{V}{Q_2}$

$\frac{C_1^n}{Q_1} = \frac{C_2^n}{Q_2}$

$C_1 = \frac{\text{Total chlorine}}{Q}$

$C_1 = \frac{32}{16} = 2 \ mg/l$

$\frac{2^1}{16} = \frac{C_2^1}{22}$

$\Rightarrow C_2 = 2\times\frac{22}{16} = 2.75 \ mg/l$

$\therefore$ Total chlorine required for 22 MLD

$= Q_2 \times C_2$

$= 22\times 2.75 = 60.5$ kg/day

14. **Ans: 50.02 m³**

Sol: % sewage kill 'η' = $(1-e^{-kt}) \times 100$

K = 0.145

$98 = (1-e^{-0.145\times t}) \times 100$

$\Rightarrow t = 26.979$ min

Contact time "t" = detention time

$= 7.0358$ min

Volume of disinfection unit $= Q \times Dt$

$$= \frac{2670}{24 \times 60} \times 26.979 = 50.02\ m^3$$

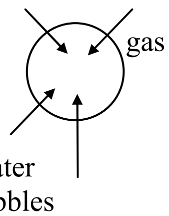
Chapter- 8
Miscellaneous Water Treatment

01. Ans: (c) **02. Ans: (d)**

03. Ans: (d) **04. Ans: (d)**

05. Ans: (b)

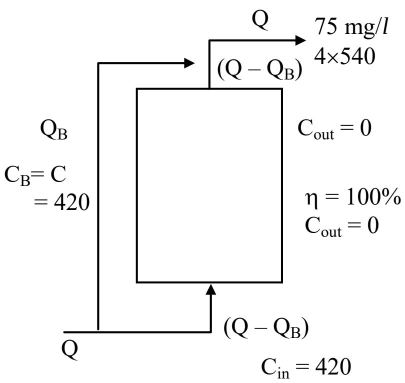
Sol: liquid –gas system absorption.



06. Ans: (b) **07. Ans: (a)**

08. Ans: 385.714

Sol:



$C_{in} = 420$ mg/*l*

$C_{out} = 75$ mg/*l*

$Q = 4 \times 540 = 2160$ lit/day

$$\eta = \frac{C_{in} - C_{out}}{C_{in}} \times 100 = \frac{420 - 75}{420} \times 100$$

$= 82.142\%$

$\eta = 100\%$

$$C_{mix} = \frac{(Q - Q_B)C_{out} + Q_B C_B}{Q}$$

$$75 = \frac{(Q - Q_B) \times 0 + Q_B \times 420}{2160} \Rightarrow Q_B$$

$= 385.714$ lit/day

Flow that can be by passed = 385.714 Lit/day

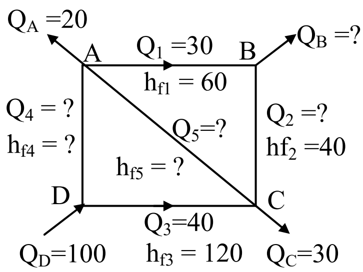
Chapter- 9
Distribution System

01. Ans: (b)

02. Ans: (d)

03.

Sol:



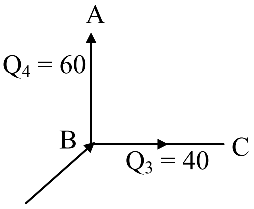
Unknown

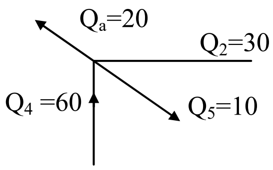
$Q_2 = ?$

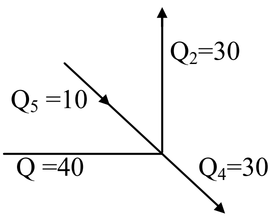
$Q_4 = 60$

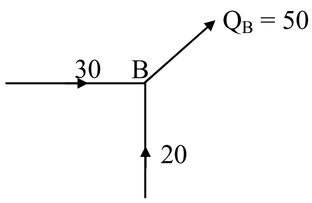
$Q_5 = 10$

$Q_B = ?$

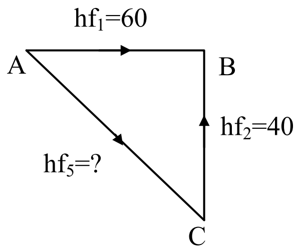








Consider loop ABCA, $\sum H_{ABCA} : 0$

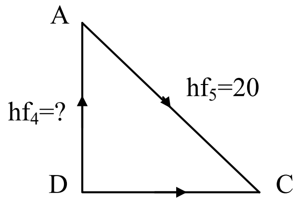


$+hf_1 - hf_2 - h\ f_5 = 0$

$+60 - 40 - hf_5 = 0$

$h_{f5} = 20$

Consider loop ACDA



$\Sigma\ H_{ABCD} = 0$

$+h_{f5} - h_{f3} + h_{f4} = 0$

$20 - 120 + h_{f4} = 0$

$h_{f4} = 100$

04. Ans: (a)

05. Ans: (a)

Waste Water Engineering

Chapter-1
Introduction to Waste Water Engineering and Estimation of DWF & WWF

01. Ans: (b)

Sol: Storm's which occurs over catchment, if the duration of storm is not given (while calculating storm water discharge resulting from the catchment) calculating time of concentration is assumed as duration of storm.

02. Ans: (a)

Sol:

$$\text{Intensity}(I) = \frac{\sum A_i \times I_i}{A}$$

$$= \frac{(40)(0.8)+(30)(0.2)+(30)(0.1)}{40+30+30}$$

$= 0.41$

$$Q_{wwF} = \frac{AIR}{360} \quad \text{Where } A = 16 \text{ ha}$$

$I = 0.41$

$R = 5$ cm/hr $= 50$ mm/hr

$$\therefore\ Q_{wwF} = \frac{(16)(0.41)(50)}{360} = 0.911 \text{m}^3/\text{sec}$$

03. Ans: 0.1736 m^3/sec, 2.015 m^3/sec

Sol: Population = 1,00,000

Q_{Dwf} = Population × percapita × factor

= 1,00,000 × 200 × 0.75

= 15 × 10^6 lpcd = 15 MLD

= 0.1736 m^3/sec

$$Q = \frac{AIR}{360} \Rightarrow R = \frac{25.4a}{t_c + b} = \frac{25.4 \times 40}{50+20}$$

= 14.51 mm/hr

$$= \frac{100 \times 0.5 \times 14.15}{360} = 2.015 \text{ m}^3/\text{sec}$$

04. Ans: 2.508 m^3/sec

Sol: P = 40000

A = 75 ha

I = 0.70

Factor = 0.70

Q_{DWF} = Population × rate of flow × factor

Q_{DWF} = 40000×120×0.70

= 0.0388 m^3/sec

$$Q_{WWF} = \frac{AIR}{360} = \frac{75 \times 0.70 \times 16.93}{360}$$

Q_{wwf} = 2.47 m^3/sec

$$R = \frac{25.4 \times 40}{40+20} = 16.93 \text{ mm/hr}$$

Combined discharge = 0.0388 + 2.470

= 2.5088 m^3/sec

05. Ans: 0.323 m^3/sec

Sol: A = 1 km^2 = 100 ha; P = 1000 no/ha

Rate of flow = 200 *l*pcd

Factor = 0.80

P = 100000

$$I = \frac{1.2\times10}{24} = 0.5\text{mm/hr}$$

$$Q_{DWF} = \frac{100000\times200\times0.80\times10^{-3}}{24\times3600}$$

$$Q_{DWF} = 0.185\,\text{m}^3/\text{sec}$$

R = 0.5 mm/h

$$Q_{WWF} = \frac{AIR}{360} = \frac{100\times1\times0.5}{360}$$

Q_{WWF} = 0.138 m^3/sec

$Q_{Design} = Q_{DWF} + Q_{WWF}$

= 0.185 + 0.138 = 0.323 m^3/sec

Chapter-2
Design of Sewers

01. Ans: (b)

02. Ans: 1.311 m

Sol: A = 150 ha

P = 50,000

V = 3.2 m/sec [$t_c = t_e + t_f$ = 5+20 =25 min]

t_e = 5 min

t_f = 20 min

Q = 270 *l*t/d/c

Impermissibility factor = 0.45

Factor = 0.75

Q_{DWF} = 50000 × 270 × 0.75

= 0.117 m^3/sec

$$Q_{WWF} = \frac{AIR}{360}$$

$$= \frac{150\times0.45\times22.57}{360}$$

$$\left[\begin{array}{l} R = \dfrac{25.4a}{t_c + b} \\ = \dfrac{25.4(40)}{25+20} = 22.57 \end{array}\right]$$

= 4.23

$Q = Q_{DWF} + Q_{WWF}$ = 0.117+4.23

= 4.34 m^3/sec

Q = AV

$$4.34 = \frac{\pi}{4}\times D^2\times3.2$$

$$D^2 = \frac{4\times4.34}{\pi\times3.2} = 1.72$$

$$D = \sqrt{1.72} = 1.311\text{m}$$

03. Ans: (c)

Sol: $V = \frac{1}{n}.(R)^{2/3}.S^{1/2}$

The velocity of flowing full and flowing half will be same

V = 1 m/sec

04. Ans: 1 in 449

Sol: dia = 45 cm = 0.45 m

Population = 30000; $Q_{design} = 3.5\, Q_{DWF}$

S = ?

Running full

n = 0.012 Factor = 0.80

Rate of water supply = 150 *l*pcd

Q_{DWF} = Population ×per capita water supply ×factor

$$Q_{DWF} = \frac{30000\times150\times0.80}{24\times10^3\times60\times60}$$

Q_{DWF} = 0.0416 m^3/sec

Q_{design} = 3.5× 0.0416

= 0.1456 m^3/sec

Running full $A = \frac{\pi}{4}D^2$

$$R = \frac{D}{4}$$

Q_{design} = A. V

$$= \frac{\pi}{4}D^2 \times \frac{1}{n}\times(R)^{2/3}\times(S)^{1/2}$$

$$0.1456 = \frac{\pi}{4}\times(0.45)^2\times\frac{1}{0.012}\times\left(\frac{0.45}{4}\right)^{2/3}\times(S)^{1/2}$$

S = 0.0022

$$S = \frac{1}{448.5} \simeq \frac{1}{449}$$

05. Ans: 1.353 m^3/sec

Sol: Q = ?

d = 1.25 m

$$S = \frac{1}{360}, n = 0.011$$

Half-full

$$A = \frac{\pi}{8}D^2 \quad R = \frac{D}{4}$$

$$Q = A.V = \frac{\pi}{8}D^2\times\frac{1}{n}(R)^{2/3}(S)^{1/2}$$

$$Q = \frac{\pi}{8}\times(1.25)^2\times\frac{1}{0.011}\times\left(\frac{1.25}{4}\right)^{2/3}\times\left(\frac{1}{360}\right)^{1/2}$$

Q = 1.35 m^3/sec

06. Ans: 0.656 m/sec, 1 in 2160

Sol: V_{self} = ?

Dia of sand particles d = 1 mm

S = 2.65

K = 0.1

f = 0.03

n = 0.013

$$V_{self} = \sqrt{\frac{8K}{f}(S-1)g.d}$$

$$V_{self} = \sqrt{\frac{8\times0.1}{0.03}(2.65-1)\times9.81\times1\times10^{-3}}$$

V_{self} = 0.656 m^3/sec

$$V_{self} = V = \frac{1}{n}.(R)^{2/3}(S)^{1/2}$$

$$0.656 = \frac{1}{0.03}.\left(\frac{1}{4}\right)^{2/3}(S)^{1/2}$$

$$S = 4.63\times10^{-4} = \frac{1}{2159}$$

S = 1 in 2159 ≃ 1 in 2160

07. Ans: 0.36 m, 0.027 m^3/sec, 0.43 m/sec

Sol: dia of sewer D = ?

n = 0.013

Q = 0.05 m^3/sec

S = 1 in 1000

Flow full

$A = \frac{\pi}{4}D^2$

$R = \frac{D}{4}$

Q = A. V

$Q = \frac{\pi}{4}D^2.\frac{1}{n}.(R)^{2/3}.(S)^{1/2}$

$0.05 = \frac{\pi}{4}D^2.\frac{1}{0.013}.\left(\frac{D}{4}\right)^{2/3}.\left(\frac{1}{1000}\right)^{1/2}$

D = 0.36 m

If the flow were at 0.60 depth

d = 0.60 D

q = ?

V = ?

$\frac{d}{D} = \frac{0.6D}{D} = 0.60$

d/D	q/Q	v/V
0.60	0.54	0.88

$\frac{q}{Q} = 0.54$

q = 0.54 × 0.05 = 0.027 m^3/sec

$\frac{v}{V} = 0.88$

$V = 0.88 \times \frac{1}{n}.(R)^{2/3}.(S)^{1/2}$

$V = 0.88 \times \frac{1}{0.014}.\left(\frac{0.36}{4}\right)^{2/3}.\left(\frac{1}{1000}\right)^{1/2}$

V = 0.43 m /sec

08. Ans: (a)

Sol: Slope = $\frac{1}{400}$

V = 0.7 m/s

dia of sewer = 600 mm

slope = $\frac{1}{200}$

V = ?

$V = \frac{1}{n}(R)^{2/3}(S)^{1/2}$ $\quad R = \frac{D}{4}$

$R = \frac{0.6}{4}$

$\frac{1}{n} = \frac{V}{(R)^{2/3}(S)^{1/2}}$

$\frac{1}{n} = \left[\frac{V}{(R)^{2/3}(S)^{1/2}}\right]_{full} = \left[\frac{V}{(R)^{2/3}(S)^{1/2}}\right]_{half}$

$= \left[\frac{0.7}{\left(\frac{0.3}{4}\right)^{2/3}\left(\frac{1}{400}\right)^{1/2}}\right]_{full} = \left[\frac{V}{\left(\frac{0.6}{4}\right)^{2/3}\left(\frac{1}{200}\right)^{1/2}}\right]$

V = 1.59 m/s

09. Ans: (c)

Sol: $\frac{Q_{full}}{Q_{half}} = \frac{A_{full}V}{A_{half}V} = \frac{\frac{\pi}{4}D^2 \times v}{\frac{\pi}{8}D^2 \times v} = 2$

The Velocity remains same for the pipe flowing full and half full, if diameter and bed slope remains same.

10. Ans: (c)

Sol: D = 300 mm ;

$q = 1728\ m^3/day$;

n = 0.015 ;

S = 1/280

To find 'Q'

$$Q = A.V = \frac{\pi}{4}.D^2 \times \frac{1}{n}.R^{2/3}.S^{1/2}$$

$$= \frac{\pi}{4}.(0.3)^2 \times \frac{1}{0.015} \times \left(\frac{0.3}{4}\right)^{2/3} \times \left(\frac{1}{280}\right)^{1/2}$$

$= 0.05\ m^3/sec$

$= 4320\ m^3/day$

$$\therefore \frac{q}{Q} = \frac{1728}{4320} = 0.4$$

For $\frac{q}{Q} = 0.4$; from the graph (given).

$$\frac{d}{D} = 0.5$$

For $\frac{d}{D} = 0.5$, the $\frac{v}{V} = 0.8$

$$\therefore \frac{d}{D} = 0.5$$

$\Rightarrow$ d = 0.5 × 0.3 = 150 mm

$$\therefore \frac{\vartheta}{V} = 0.8$$

$\vartheta = 0.8 \times 0.708 = 0.57$ m/sec

Chapter-3
Characteristics of Sewage

01. Ans: 212.19 mg/lit

Sol: Ultimate BOD L_o = ?

6 ml → Waste → $D_o = 0$

294 ml → distilled → $D_0 = 8.6$ mg/l

$(D_0)_{Final} = 5.4$ mg/l

K(base) e = $0.25d^{-1}$

$$(D_0)_{Initial} = \frac{V_D(D_0) + V_S(D_0)_S}{V_0 + V_S}$$

$$= \frac{294 \times 8.6 + 6 \times 0}{294 + 6}$$

$(D_0)_I = 8.428$ mg/l

$y_5^{20^0C} = [(D_0)_I - (D_0)_F] \times D.F$

$$y_5^{20^0C} = [8.428 - 5.4] \times \frac{300}{6}$$

= 151.4 mg/l

5 day BOD at 20^0C $y_5^{20^0C}$

$$y_t^{TOC} = L_0\left(1 - e^{K_t \times t}\right)$$

$$y_t^{20^0C} = 151.4 = L_0\left(1 - e^{K_{20} \times 0}\right)$$

$151.4 = L_0(1 - e^{-0.25 \times 5})$

L_0 = Ultimate BOD 'L_0' = 212.19 mg/l

02. Ans: (d)

Sol: 5% dilution of sample = $\frac{100}{5} = 20$

$$= \frac{300}{15} = 20$$

$(D_0)_F = 3.80$ mg/l , $(D_0)_{blank} = 8.80$ mg/l

$(D_0)_S = 0.8$ mg/l

$$(D_0)_I = (D_0)_{mixer} = \frac{V_0(D_0)_0 + V_S(D_0)_S}{V_0 + V_S}$$

$$= \frac{285 \times 8.8 + 0.8}{285 + 15} = 8.4 \text{ mg}/l$$

$$y_5^{20} = [(D_0)_I - (D_0)_F]DF$$

$$= (8.4 - 3.80) \times 20 = 92 \text{ mg}/l$$

03. Ans: (c)

Sol: Fail in finding the BOD of waste water

04. Ans: 90 mg/*l*

Sol: $y_5^{30^0C}$ = sewage sample = 110 mg/l,

$K_{D(20)}$ = 0.1/day = base 10

$$y_5^{30^0C} = L_0\left(1 - e^{-K_t t}\right)$$

$$y_5^{20^0C} = ?$$

K_{20} (base 10) = 0.1 d^{-1} = 2.3 × 0.1

= 0.23d^{-1}

$$L_0 = \frac{y_t^{T^0C}}{\left(1 - e^{-k_t t}\right)} = \frac{y_5^{30^0C}}{(1 - e^{-k_t t})}$$

$$y_5^{30} = L_0\left(1 - e^{-K_{30} \times t}\right)$$

$K_T = K_{20}(1.047)^{T-20}$

$K_{30} = 0.23(1.047)^{30-20}$

= 0.364d^{-1}

$110 = L_0(1-e^{-0.364\times5})$

$L_0 = 131.78$ mg/l

$$y_5^{20} = L_0\left(1 - e^{-K_{20} \times 5}\right)$$

$131.78(1-e^{-0.23\times5}) = 90$ mg/l

05. Ans: 246.36 mg/*l*

Sol: $y_1^{30^0C}$ = 110 mg/l, $y_5^{20^0C}$ = ?

$K_{(20)} = 0.1d^{-1}$

= 2.3 × 0.1

= 0.23d^{-1}

$$y_1^{30^0C} = L_0\left(1 - e^{-0.364\times1}\right)$$

$110 = L_0(1-e^{-364.\times1})$

$L_0 = 360.5$ mg/l

$K_{30} = 0.23(1.047)^{30-20}$

= 0.364d^{-1}

$$y_5^{20} = L_0(1 - e^{-K_{20}\times0})$$

$= 360.5(1-e^{-0.23\times5}) = 246.36$ mg/l

06. Ans: 304000

Sol: $Q = 80 \times 10^6$ l/d, $y_5 = 285$ mg/l,

compute daily 5 day O_2 demand

Total strength of waste = Q × y

= 80 × 285

= 22800 kg/day

$$\text{Population equation} = \frac{Q \times y}{\text{per capita BOD}}$$

$$75 \text{ g} = \frac{22800}{\text{per capita BOD}}$$

$$\text{Per} = \frac{22800}{75 \times 10^{-3}} = \frac{\text{kg/day}}{\text{kg/day}}$$

Population equation = 304000 persons

07. Ans: 93.72%

Sol: $Sr = 100(1\text{-}0.794^{t})$

$= 100(1\text{-}0.794^{12})$

$Sr = 93.72\%$

08. Ans: 1%

Sol: $BOD_5 = 600$ mg/l,

K = 0.23/d (base e); $K = 0.23d^{-1}$, $L_0 = ?$

BOD_u remain unaxidised after 20 days =?

$y_5^{20^0C} = L_0\left(1 - e^{K_1 t}\right)$

$600 = L_0(1 - e^{-0.23\times5})$

$L_0 = 878.01$ mg/l

$L_{20} = L_0 e^{-Kt}$

$= 878.01 \times e^{Kt}$

$= 878.01 \times e^{-0.23\times20} = 8.82$ mg/l

% of unoxidised = $\frac{8.82}{878.01}\times100 = 1\%$

1% of BOD after 20 days

09. Ans: (a)

Sol:

Waste water Volume ml	Initial D_0 mg/l	D_0 after 50 day mg/l
5	9.2	6.9
10	9.1	4.4
50	8.4	0.0

Diluted = $\frac{300}{5} = 60$

$y_5^{20} = [(D_0)_I - (D_0)_F] \times DF$

$= [9.2 - 6.9]\times 60$

$= 138$ mg/l

$y_5^{20} = [9.1 - 4.4]\times\frac{300}{10} = 141$ mg/l

$BOD_{avg} = \frac{138+141}{2} = 139.5$ mg/l

10. Ans: (c)

Sol: $K = 0.01h^{-1}$ (base)

$= 0.01 \times 24\ h^{-1}$

$= 0.24\ h^{-1}$

$y_5^{20} = L_0(1 - e^{-Kt})$

$190 = L_0(1 - e^{-0.0124\times5})$

$L_0 = \frac{190}{\left(1 - e^{0.24\times5}\right)} = 271.89$ mg/l

11. Ans: (d)

Sol: $y_5^{20} = 180$ mg/l, $K_T = K_{20}(1.047)^{T-20}$

BOD = 2.5 day

$\left[y_5^{20} = L_0\left(1 - e^{-0.18\times5}\right)\right]$

$y_{2.5}^{T^0C} = 180$ mg/l

$y_5^{20} = y_{2.5}^{T^0C}$

$L_0(1 - e^{-K_{20}\times5}) = L_o(1 - e^{-K_T\times2.5})$

$K_{20} \times 5 = K_{20}(1.047)^{T-20}\times2.5$

$(1.047)^{T-20} = \frac{5}{2.5}$

$(T-20)Ln(1.047) = Ln2$

$(T-20)0.045 = 0.693$

$T = 35^0C$

12. Ans: (c)

Sol: $BOD_3 = 75$ mg/*l*, K = 0.345d^{-1} (base e)

BOD = 10 days = ?

$y_3^{20} = L_0(1-e^{-0.345\times3})$

$L_0 = 116.31$ mg/*l*

$L_t = 116.31(1-e^{-0.345\times10})$

$L_t = 112.61$ mg/*l*

$L_0 - L_t = 116.31 - 112.61 = 3.7$ mg/*l*

13. Ans: (b) **14. Ans: (b)**

15. Ans: (c)

16. Ans: (b)

Sol: $y_5^{20^\circ C} = [(DO)_I - (DO)_F]\times DF$

$= [8.5-5.5]\times\frac{100}{2} = 150$ mg/*l*

17. Ans: (a)

18. Ans: (c)

Sol: y = 162 mg/*l*

Q = 1000 m^3/day

Q = 1000 × 1000 *l*pcd = 1 MLD

Per capita BOD = 80 gm/capita

Population equivalent

$= \frac{\text{total BOD}}{\text{per capita BOD}}$

$= \frac{Q\,y}{80\times10^{-3}} = \frac{1\times162}{80\times10^{-3}} = 2025$

19. Ans: (b)

Sol:

$[(D_0)]_I = 8$ mg/*l* , $(D_0)_f = 2$mg/*l*

Dilution factor = $\frac{300}{2}$ = 150 m*l*

5 days BOD = $[(D_0)_I - (D_F)]\times D.F$

= (8–2) × 150

= 900 mg/*l*

20. Ans: (a)

21. Ans: (c)

Sol:

y_5^{20} = 250 mg/*l*, T = 30^{0}C, t = ?

y_5^{30} = 250 mg/*l*

$L_0 = \frac{250}{(1-e^{-K_t\times t})}$

$K_{30} = K_{20}(1.047)^{30-20}$

$y_5^{20} = y_5^{30}$

$L_0(1-e^{K_{20}\times5}) = L_0(1-e^{-K_{30}\times t})$

$K_{20}\times5 = K_{30}\times t$

$K_{20}\times5 = K_{20}(1.047)^{30-20}\times t$

$t = \frac{5}{(1.047)^{10}} = 3.158 = 3.3$ days

22. Ans: (a)

23. Ans: 128.1 mg/*l*

Sol:

$y_5^{20^\circ} = ?$

$y_7^{20^\circ} = 150 \text{mg}/\ell$

$K = 0.23d^{-1}$

$150 = L_o (1 - e^{-23\times7})$

$L_o = 187.47$ mg/*l*

$y_5^{20^\circ C} = L_0\left(1 - e^{-0.23\times5}\right)$

$y_5^{20^\circ C} = 187.47\left(1 - e^{-0.23\times5}\right)$

$y_5^{20^\circ} = 128.11 \text{mg}/\ell$

Chapter- 4
Treatment of Sewage

01. Ans: (a)

02. Ans: (d)

Sol: 12 m × 1.50, H = 0.8 m, Q = 720 m^3/hr

L = 12 m, B = 1.50 m

$$\text{Surface loading rate} = \frac{Q}{\text{surface area}}$$

$$= \frac{720}{12\times1.50} = 40 \text{ m}^3/\text{hr/m}^2$$

$$= 40000 \text{ lit/hr/m}^2$$

$$D.T = \frac{\text{Volume of G.C}}{Q} = \frac{L\times B\times H}{Q}$$

$$= \frac{12\times1.5\times0.8}{\frac{720}{60}} = 1.2 \text{ min}$$

03. Ans: (d)

04. Ans: (b)

Sol: Q =3 m^3/sec, cross section of grit chamber = ?

$$\text{Cross section area (B.H)} = \frac{Q}{V_H}$$

$$A = \frac{Q}{V_H}$$

$$A = \frac{3}{0.3} = 10 \text{ m}^2$$

05. Ans: (b)

Sol: G = 2.70

d = 0.21 mm

V_s = ?

$\upsilon = 1\times10^{-2}$ cm^2/sec = 1×10^{-6} m^2/sec

For laminar flow condition

$$V_s = \frac{g(S-1)d^2}{18\upsilon}$$

$$V_s = \frac{9.81\times(2.70-1)\times(0.21\times10^{-3})^2}{18\times1\times10^{-6}}$$

V_s = 0.04089 m/sec

V_s = 4.089 cm/sec

06. Ans: (b)

Sol: L = 7.5 m

V_H = 0.3 m/sec

H = 0.9m

$$V = \frac{\mu}{\rho} = \frac{1.002\times10^{-3}}{1000}$$

$= 1.002\times10^{-6}$ m^2/sec

For η = 100%, $\frac{L}{V_H} = \frac{H}{V_S}$

$\Rightarrow V_S = 0.036$ m/sec

$$V_S = \frac{g}{18}\left((s-1).\frac{d^2}{v}\right)$$

$$0.036 = \frac{9.8}{18}(2.5-1).\frac{d^2}{(1.002\times10^{-6})}$$

d = 0.21 mm

07. Ans: (d)

Sol: Q = 5005 m^3/d

$V_0 = 35\ m^3/m^2/d$

$$\text{Surface area} = \frac{Q}{V_0} = \frac{5005}{35}$$

$$\frac{\pi}{4}d^2 = 143$$

d = 13.5 m

Chapter- 5
Activated Sludge Process

01. Ans: (a)

Sol: y_i = 180 mg/l, V_L = 550 gm of BOD per 1 cu.m of volume

Q = 50 × 10^6 l/d = 50 MLD, V = ?

$$\text{V.L.R.} = \frac{Qy_i}{V}$$

$$550\times10^{-3} = \frac{50\times180}{V}$$

$$550\times10^{-3}\times V = 9000$$

$$V = \frac{9000}{550\times10^{-3}}$$

V = 16363 m^3

02. Ans: (d)

Sol: Q = 50 ×10^6 l/d = 50 MLD, y_i = 180 mg/l,

$\frac{F}{M}$ = 0.5 d^{-1}, X = 1800 mg/l, V = ?

$$\frac{F}{M} = \frac{Qy_i}{VX}$$

$$0.5 = \frac{50{,}000\times180}{1800\times V}$$

$$\Rightarrow V = \frac{50000\times180}{0.5\times1880}$$

V = 10,000m^3

03. Ans: (a)

Sol: $X = 2000$ mg/l

$$SVI = \frac{V}{X} = \frac{\text{Volume occupied in ml}}{\text{MLSS in gm}}$$

$$= \frac{176}{\frac{2000}{10^3}} = \frac{176}{2} = 88 \text{ ml/gm}$$

Common Data for Question Nos. 4 & 7

04. Ans: (c)

Sol: $Q = 35{,}000$ m^3/d, $V = 10900$ m^3,

$y_i = 250$ mg/l, $y_e = 20$ mg/l

MLSS = 2500 mg/l

$$\text{Aeration period} = \frac{V}{Q} = \frac{10900}{\frac{35000}{24}} = 7.47 \text{ hrs}$$

05. Ans: (b)

Sol: $$\frac{F}{M} = \frac{Q(y_i)}{VX}$$

$$= \frac{35000(250)}{10{,}900 \times 2500} \ \frac{m^3/d/(mg)/l}{m^3/mg/l}$$

$$= 0.32 \text{ d}^{-1}$$

06. Ans: (c)

Sol: $$\eta_{BOD} = \frac{y_i - y_e}{y_i} \times 100$$

$$= \frac{250-20}{250} \times 100 = 92\%$$

07. Ans: (a)

Sol: $$\text{Sludge age } \theta_c = \frac{VX}{Q_w Y_U + Q_e \times X_e}$$

$$= \frac{10900 \times 2500}{220 \times 9700 + (35000 - 220)30} \ m^3 \times \frac{mg}{l}$$

=8.57 days

08. Ans: (c)

Sol: SVI = 88 ml/gm, X_u = ?

$$X_u = \frac{10^6}{SVI} = \frac{10^6}{88} = 11364 \text{ mg}/l$$

09. Ans: (d)

Sol: $V = 400$m^3, X= 1000 mg/l

Total amount of MLSS in aeration

Tank = VX

= m^3 (mg/l)

= $400 \times 1000 \times (1000 \times 10^{-6})$

= 400 kg

10. Ans: (d)

11. Ans: (c)

Sol: X = 2800 mg/l carried out 1 lit sample

V = 200 ml

$$SVI = \frac{200}{2.8} = 71.4 \text{ m}l/\text{gm}$$

Common data for Q 12 & 13

12. Ans: (c)

Sol: Given,

$Q = 500\ m^3/h$

$y_i = 150\ mg/l$

$y_e = 10\ mg/l$

D T = 8 hours

$Q_C = 240$ hours

$V = 4000\ m^3$

$X = 2000\ mg/l$

$$\frac{F}{M} = \frac{Q.y_i}{VX} = \frac{500 \times 150}{4000 \times 2000}$$

$$\frac{F}{M} = 9.375 \times 10^{-3} \text{ per hour}$$

$$\frac{F}{M} = 9.377 \times 10^{-3} \times 24 \text{ per day}$$

$$\frac{F}{M} = 0.225 \text{ per days}$$

13. Ans: (c)

Sol: $$Q_C = \frac{\text{mass of solid reactors}}{\text{mass of solid wasted per day}}$$

$$Q_C = \frac{VX}{\text{mass of solid wasted per day}}$$

$$\text{Mass of solid wasted/day} = \frac{VX}{Q_C}$$

$$= \frac{4000 \times 2000}{\frac{240}{24}} \times 1000 \times \frac{1}{10^6}$$

= 800 kg/day

14. Ans: (d)

Sol: $X = 4000\ mg/l = 4\ gm/l$

Volume = 200 ml

$$SVI = \frac{\text{Volume}}{X} = \frac{200}{4}$$

SVI = 50 ml/gm

15. Ans: (a)

16. Ans: i. ($2000\ m^3$) **ii. (4.8 hrs)**

iii. (86.66%) **iv. $0.75\ kg/day/m^3$**

v. 600 kg/day **vi. $60\ m^3/day$**

vii. 100

viii. 0.428

ix. $4285.71\ m^3/day$

Sol: Given:

Flow rate (Q) = $10,000\ m^3/day$

Inflow BOD (y_i) = 150 mg/lit

Outflow BOD (y_E) = 20 mg/lit

MLSS (X) = 3,000 mg/lit

We know

$$\frac{F}{M} = \frac{Qy_i}{VX}$$

$$VX = \frac{10{,}000 \times 150}{0.25}$$

$$VX = 6 \times 10^6$$

(I) Volume of aeration tank

$$V = \frac{6 \times 10^6}{3 \times 10^3} \quad (\because X = 3000\ mg/lit)$$

$\therefore$ Volume = $2000\ m^3$

(II) Aeration period = $\frac{2000}{10{,}000} \times 24 = 4.8$ hours

(III) B.O.D removal efficiency

$$= \frac{y_i - y_E}{y_i} \times 100$$

$$= \frac{150-20}{150} \times 100$$

η = 86.66%

(IV) Volumetric loading rate

$$V_L = \frac{Qy_i}{V} = \frac{10{,}000 \times 150}{2000} \times \frac{1000}{10^6}$$

$$= 0.75 \text{ kg/day/m}^3$$

(V) Mass of sludge wasted per day

$$\text{MCRT} = \theta_c = \frac{VX}{Q_w x_u + Q_e x_e}$$

$$10 = \frac{200\text{m}^3 \times 300\text{mg/L} \times \frac{100\text{L}}{\text{m}^3} \times \frac{1\text{kg}}{10^9\text{mg}}}{\text{mass of sludge wasted per day}}$$

$\therefore$ Mass of sludge wasted per day = 600 kg/day

(VI) Volume of sludge wasted per day

$Q_w x_u + Q_e x_e$ = mass of sludge wasted per day

$Q_w(10{,}000) + (Q - Q_w)(0) = 600$ kg/day

$$Q_w = \frac{600}{10{,}000} = 0.06\,\text{MLD}$$

$$Q_w = 0.06 \times \frac{10^6\text{L}}{1\text{mL}} \times \frac{1\text{m}^3}{1000\text{L}}$$

$Q_w = 60 \text{ m}^3\text{/day}$

(VII) $$\text{S.V.I} = \frac{10^6}{X_w} = \frac{10^6}{10{,}000} = 100 \text{ ml/gm}$$

S.V.I = 100

(VIII) $$\frac{Q_R}{Q} = \text{Recycling ratio} = \frac{X}{X_w - X}$$

$$\Rightarrow \frac{Q_R}{Q} = \frac{3000}{10000-3000} = 0.4285$$

(IX) $\therefore$ Rate of return sludge

$Q_R = 0.4285 \times 10{,}000$

$Q_R = 4285 \text{ m}^3\text{/day}$

Chapter- 6
Trickling Filters

01. Ans: (c)

Sol: $y_i = 200$ mg/*l*

$y_e = 40$ mg/*l*

$$\eta = \frac{y_i - y_e}{y_i} \times 100$$

$$\eta = \frac{200-40}{200} \times 100 = 80\%$$

02. Ans: (b)

Sol: OLR = 0.175 kg/m^3/day

$y_i = 150$ mg/*l*

$y_e = ?$

$$\eta = \frac{100}{1 + 0.0044\sqrt{\frac{Q.y_i}{V.F}}}$$

1 ha.m = 10^4 m^3

1 m^3 = 10^{-4} ha. m

$\frac{Qy_i}{V}$ kg of BOD/day/ ha.m

OLR = 0.175 $\times 10^4$ kg/ha.m/day

$$\eta = \frac{150 - y_e}{150} \times 100$$

$$\frac{150 - y_e}{150} \times 100 = \frac{100}{1 + 0.0044\sqrt{\frac{0.175 \times 10^4}{1}}}$$

y_e = 23.31 mg/*l*

03. Ans: 2.14m, 5142.85m³, 84.45%, 23.3 mg/*l*

Sol: Q = 6 MLD

y_i = 150 mg/*l*

OLR = 175 gm/m³/day

Depth of TF = ?

Vol. of TF = ?

SLR = 2500 *l*/m²/day

η = ?

y_e = ?

$$\text{Surface loading rate} = \frac{Q}{\text{surface area}}$$

$$\frac{\pi}{4} \times d^2 = \frac{6 \times 10^6}{2500}$$

d = 55.27 m

$$\text{organic loading length} = \frac{Q.y_i}{V}$$

$$V = \frac{150 \times 6}{175 \times 10^{-3}}$$

V = 5142.85 m³

$$\text{Depth of TF} = \frac{\text{Vol. of TF}}{\text{surface area of TF}}$$

$$d = \frac{5142.85}{2400}$$

d = 2.14 m

OLR = 175 $\times 10^{-3} \times 10^4$ kg/ha. m/day

OLR = 1750 kg/ha. m/day

$$\eta = \frac{100}{1 + 0.0044\sqrt{\frac{Qy_i}{VF}}}$$

$$\eta = \frac{100}{1 + 0.0044\sqrt{175 \times 10^{-3} \times 10^4}}$$

η = 84.45%

$$\eta = \frac{y_i - y_e}{y_i} \times 100$$

$$84.45 = \frac{150 - y_e}{150} \times 100$$

y_e = 23.32 mg/*l*

04. Ans: 633 m³

Sol: Single stage TF; y_e = 20 mg/*l*

y_i = 120 mg/*l*; Q = 2200 m³/day

R = 4000 m³/day; V = ?

∵ Recirculation is there it is high rate TF

$$\frac{R}{I} = \frac{4000}{2200} = 1.81$$

$$F = \frac{1 + \frac{R}{I}}{\left(1 + 0.1\frac{R}{I}\right)^2}$$

F = 2.017

Q = 2200 m³/day

Q = 2200 ×1000 *l*pcd

Q = 2.2 MLD

$$\eta = \frac{y_i - y_e}{y_i} \times 100$$

$$\eta = \frac{120-20}{120} \times 100 = 83.33\%$$

$$83.33 = \frac{100}{1+0.0044\sqrt{\frac{2.2\times120}{V\times2.017}}}$$

V = 0.0633 ha. m

V = 633 m^3

05. Ans: (b) **06. Ans: (b)**

Chapter- 7
Sludge Digestion

01. Ans: 17105.62 kg/day, 16.68m^3/day

Sol: Q = 4.5 MLD

Total dry solids = Q ×sewage containing

= 4.5 × 275= 1237.5 kg

Mass of sludge produced = $\frac{100}{(100-P_1)}M$

Solids concentration in sludge

$$= \frac{55}{100} \times 275$$

= 151.25 mg/*l*

Total mass of dry solids produced/day

$$= \frac{100}{(100-\eta_C)} \times M$$

$$= \frac{100}{(100-96)} \times 680.625$$

= 17015.625 kg/day

(ii) $\frac{\rho_{sludge}}{\rho_w} = S_{sludge}$

$$1.02 = \frac{\rho_{Sludge}}{\rho_w}$$

ρ_{sludge} = 1.02 ×1000 = 1020 kg/m^3

$$\rho_{sludge} = \frac{\text{mass of sludge}}{\text{Volume of sludge}}$$

$$1020 = \frac{17015.625}{\text{Volume of sludge}}$$

$$\text{Volume of sludge} = \frac{17015.625}{1020}$$

= 16.68 m^3/day

02. Ans: (d)

03. Ans: (c)

Sol: $\frac{V_2}{V_1} = \frac{100-P_1}{100-P_2}$

$$= \frac{100-99}{100-96} = 25\%$$

% of reduction is volume = 100 V_1–25V_1

= 75% V_1

04. Ans: (b)

05. Ans: (d)

Sol: $V_2 = \frac{100-P_1}{100-P_2} \times P = \frac{100-98}{100-96} \times P$

$$= \frac{1}{2}(P) = \frac{P}{2}$$

06. Ans: (c)

07. Ans: 1.011, 1011 kg/m^3

Sol: Given solids content = 2%

Let ρ_{solid} be the mass density of solids

Solids again contain 70% volatile & 30% of non-volatile

$$\therefore \text{we know} = \frac{100}{S_{solids}} = \frac{70}{S_{volatile}} + \frac{30}{S_{non-volatile}}$$

$$\Rightarrow \frac{100}{S_{solids}} = \frac{70}{2.2} + \frac{30}{2.7}$$

$$\Rightarrow S_{solids} = 2.329 \simeq 2.4$$

Now let mass density of sludge as ρ_s and specific gravity as S_s

$$\therefore \frac{100}{S_s} = \frac{2}{S_{solids}} + \frac{98}{S_{water}} \Rightarrow \frac{100}{S_s} = \frac{2}{2.4} + \frac{98}{1}$$

$$\Rightarrow S_s = 1.011$$

$$\rho_s = 1011 \text{ kg/m}^3$$

Chapter- 8
Septic Tanks

01. Ans: 6.12 m × 3.16 m

Sol: Discharge = 150 × 200 × 10^{-3} = 30 m^3/day

Given detention time = 24 hours

$\therefore$ Volume of septic tank (v) = t×Q

$$= \frac{(24)(30)}{24}$$

$$= 30 \text{ m}^3$$

$$\therefore \text{Area of septic tank} = \frac{(30)}{(1.5)} = 20 \text{ m}^2$$

$$\therefore \text{Given } \frac{L}{B} = 2:1$$

L × B = 20

2B × B = 20

B = 3.16 m

and L = 6.12 m

02. Ans: 0.6 m

Sol: Volume of sludge produced

$= 30\times10^{-3}\times200\times2$

$= 12 \text{ m}^3$

$\therefore$ depth of sludge zone

$$= \frac{\text{Volume of sludge}}{\text{Area of septic tank}}$$

$$= \frac{12}{20} = 0.6\text{m}$$

03. (a) 9.6 m^3 (b). 12 m^3

Sol: Given sewage flow = 150 *l*pcd

Sewage discharge = 150 × 10^{-3}×120 m^3/day

= 18 m^3/day

$$\text{Detention period} = \frac{V}{Q} = \frac{4\times2\times1.5}{18\text{m}^3/\text{day}} = 16\,\text{hrs}$$

(a) Volume of sludge = (Q × detention time)

$= 40 \times 10^{-3}\times120\times2$

$V = 9.6 \text{ m}^3$

(b) Volume of soakpit

$$= \frac{Q}{\text{Percolation capacity}}$$

$$= \frac{(150\times120)}{(1500)} = 12\ m^3$$

Chapter- 9
Oxidation Ponds

01. Ans: L = 282.84m, B = 70.71m, D.T. = 50 days

Sol: Q = population × per capita, supply×Factor

= 10,000× 100×0.8

$$= \frac{1000000\times0.8}{10^6} = 0.80\text{MLD}$$

y = 40 g/day

D.T = ?

80% of BOD removal BOD loading rate

= 200 kg = BOD/hect/d

$$\text{Surface area of pond} = \frac{Q\times y_i}{\text{BOD loading rate}}$$

$$= \frac{0.8\times500}{200} = 2\ \text{ha}$$

Total BOD = Q $\times y_i$

= population × per capita BOD

= 0.8 × y_i = 10,000 × 10040 × 10^{-3}

y_i = 500 mg/li

Surface area = 2 ha = 2 × 10^4 m^2

L = 4B

L × B = 2 × 10^4

4B × B = 2 × 10^4

$$B = \sqrt{\frac{2\times10^4}{4}} = 70.71\ \text{m}$$

L = 4 × 70.71 = 282.84 m

$$\text{D.T} = \frac{\text{Volume of pond}}{Q} = \frac{L\times B\times H}{Q}$$

$$= \frac{282.84\times70.71\times2}{\frac{0.8\times10^6}{10^3}} = 50\ \text{days}$$

$$\eta = 80 : \frac{y_i - y_e}{y_i}\times100$$

$$80 = \frac{500 - y_e}{500}\times100$$

$$0.8 = \frac{500.y_e}{500} \Rightarrow 400 = 500 - y_e$$

y_e = 500 – 400 = 100 mg/l

02. Ans: L = 244.9 m, B = 61.23m, D = 1

Sol: P = 10,000 sewage flow = 150 $l$pcd

$$Q = \frac{10000\times150}{10^6} = 1.5\,\text{MLD}$$

y_i = 300 mg/l, y_e = 30 mg/l,

OLR = 300 kg/ha/d

k_D = 0.23d^{-1}, L:B = 4:1

$$\therefore \text{D.T.} = \frac{1}{0.23}\times\ell_n\frac{300}{30} = 10\ \text{days}$$

$$\therefore \text{Surface area} = \frac{Qy_i}{\text{OLR}} = \frac{1.5\times300}{300} = 1.5\text{ha}$$

L×B = 1.5×10^4 m^2

$4B\times B = 1.5\times10^4$

$B = 61.23$ m

$\therefore L = 4\times61.23 = 244.92$ m

Vol of oxidation pond = Q ×DT

$$= \frac{1.5\times10^6}{10^3}\times10 = 15000\,m^3$$

$$\text{Depth of pond, } H = \frac{\text{Vol. of pond}}{\text{Surface area of pond}}$$

$$= \frac{15000}{15000} = 1m$$

03. Ans: (d)

04. Ans: (c)

Sol: Q = 10000 × 200, y = 300 mg/l

Organic loading = 310 kg/day/m

Q = 2 MLD

$$\text{Surface area of pond} = \frac{Qy_i}{\text{BOD loading rate}}$$

$$= \frac{2\times300}{310} = 1.93 \backsimeq 2 \text{ ha}$$

Chapter- 10
Disposal of Sewage Effluents

01. Ans: (b)

Sol: $y_R = 200$ mg/l, $Q_R = 50\ m^3/s$

$y_w = 8$ mg/l, $Q_w = 500\ m^3/s$

$$y_{mix} = \frac{Q_R y_R + Q_w y_w}{Q_R + Q_w}$$

$$= \frac{50\times200+500\times8}{50+500}$$

$y_{mix} = 25.45$ mg/l

02. Ans: (c)

Sol: Waste water $(DO)_w = 2$ mg/l

$Q_W = 1.10\ m^3/sec$

$(DO)_R = 8.3$ mg/l

$Q_R = 8.70\ m^3/sec$

$$(DO)_{mix} = \frac{(DO)_w.Q_w + (DO)_R.Q_R}{Q_w + Q_R}$$

$$(DO)_{mix} = \frac{2\times1.10+8.3\times8.70}{1.10+8.70}$$

$(DO)_{mix} = 7.6$ mg/l

03. Ans: 13.85 mg/*l*, 20.27 mg/*l* , 5.85 mg/*l*

Sol: $Q_{Rw} = 12000\ m^3/d$, temp = 20^0C,

$y_w = 50$ mg/l

D.O = concentration = 2 mg/l

$Q_R = 40{,}000\ m^3/d$, $y_R = 3$ mg/l,

D.O = 7 mg/l, temp = 20^0C

K = 0.23 (to the base in decay curve)

$$y_{mix} = \frac{Q_R y_R + Q_w y_w}{Q_R + Q_w}$$

$$= \frac{12000\times 50 + 40{,}000\times 3}{12{,}000 + 40{,}000} = 13.84 \text{mg/lit}$$

$$(DO)_{mix} = \frac{Q_R (DO)_R + Q_w (DO)_w}{Q_R + Q_w}$$

$$= \frac{12{,}000\times 2 + 40{,}000\times 7}{12{,}000 + 40{,}000}$$

$= 5.85$ mg/l

$$y_{mix} = L_0(1-e^{-k\times t})$$

$$13.85 = L_0(1-e^{-0.23\times 5}) = \frac{13.85}{\left(1-e^{-0.23\times 5}\right)}$$

$L_0 = 20.27$ mg/l

04. Ans: (c)

Sol:

River	Waste water stream
$Q_R = 12\ m^3/sec$	$Q_w = 2\ m^3/sec$
$(L_0)_R = 5$ mg/l	$(L_0)_w = 90$ mg/l

$$(L_0)_{mix} = \frac{Q_R (L_o)_R + Q_W (L_o)_W}{Q_R + Q_w}$$

$$= \frac{12\times 5 + 2\times 90}{12+2} = 17.142 \text{ mg}/l$$

$Q_{mix} = Q_R + Q_w = 12+2 = 14\ m^3/sec$

$$\text{Velocity} = \frac{Q_{min}}{c/s\,\text{area}} = \frac{14}{50} = 0.28\,\text{m/sec}$$

Time taken by the river to travel 10 km

$$= \frac{10\times 1000}{0.28}\times\frac{1}{24\times 50\times 60}$$

T = 0.413 days

Ultimate BOD at 10 km d/s = $L_t = L_o\,e^{-kt}$

k = 0.25 1/day

$L_{0.413} = 17.142\ e^{-0.25\times 0.413} = 15.459$

$\simeq 15.46$ mg/l

Chapter- 11
Solid Waste Management

01. Ans: 3870

Sol: Total energy as discarded

$$= \frac{\Sigma P_i E_i}{100} = \frac{P_1E_1 + P_2E_2 + \ldots\ldots\ldots + P_nE_n}{100}$$

$$= \frac{[20\times 2500 + 10\times 10000 + 10\times 8000 + 10\times 14000 + 40\times 3500 + 5\times 14000 + 5\times 100]}{100}$$

= 5805 kJ/kg

Moisture content of MSW

$$= \frac{20\times 70 + 10\times 4 + 10\times 4 + 10\times 1 + 40\times 60 + 5\times 20\times 2 + 5\times 2}{100}$$

= 40%

Total energy on dry basis

$$= \text{energy as discarded}\times\frac{100}{100-\%mc}$$

$$= 5805\times\frac{100}{100-40} = 9675\,\text{kJ/kg}$$

Difference in energy content = 9675 – 5805

= 3870 kJ/kg

02. Ans: (b)

Sol: $\frac{100}{\rho Msw} = \frac{\%F.w}{\rho Fw} + \frac{\%DA}{\rho DA} + \frac{\%pla}{\rho p} + \frac{\%WS}{\rho WS}$

$= \frac{50}{300} + \frac{30}{500} + \frac{10}{65} + \frac{10}{125}$

$= \frac{100}{\rho Msw} = 0.46 \Rightarrow \frac{100}{0.46}$

$\rho Msw = 217.1\ kg/m^3$

03. Ans: (d)

04. Ans: (d)

Sol: 50 g of Co_2, 25 g = CH_4, 1 million people, rate of 500 ton/day

120 parts of MSW release 50 parts of Co_2 and 25 parts CH_4 1 part of MSW release

$= \frac{75}{120} = 0.625$ parts of green house

500 t of Msw release = 0.625 × 500

Total green house = 321.5 of green house

Per capita green house gas contribution

$= \frac{\text{house gas}}{\text{population of community}}$

$= \frac{312.5}{10,00,000} \times 1000 \times 1000$

= 312.5 mg/l

05. Ans: (b) **06. Ans: (d)**

07. Ans: (d) **08. Ans: (c)**

09. Ans: (d)

Sol: 40 + 35 = 75t

10. Ans: (b)

11. Ans: 1.46 ha

Sol: Population = 65000

Rate of solid waste = 2 kg/capita/day

Solid waste generated = 65000 ×2

= 130000 kg/day

Solid waste generator per annum

= 130000×365

= 47450000

$\text{Volume} = \frac{\text{mass}}{\text{density}}$

$V = \frac{47450000}{650}$

$V = 73000\ m^3$

$\text{Area} = \frac{\text{Volume}}{\text{depth}} = \frac{73000}{5}$

$A = 14600\ m^2 = 1.46$ ha

12. Ans: (a)

Sol: $\frac{100}{\rho_{sludge}} = \frac{c_1}{\rho_1} + \frac{c_2}{\rho_2}$

$\Rightarrow \rho_{sludge} = \frac{100}{\frac{c_1}{\rho_1} + \frac{c_2}{\rho_2}}$

13. Ans: 13.6875

Sol: Solid waste generated = $2 \times 10^5 \times 25 \times 365 \times 2$

$= 3.65 \times 10^9$ kg

Volume of un-compacted

$$S_w = \frac{3.65\times10^9}{100} = 36.5 \times 10^6 \text{ m}^3$$

Volume of compacted solid waste

$$= \frac{36.5\times10^6}{4} = 9.125 \times 10^6 \text{ m}^3$$

$$\frac{\text{Compacted fill}}{\text{Compacted solid waste}} = 1.5$$

Volume of land fill (compacted fill)

$= 9.125 \times 10^6 \times 1.5$

$= 13.6875 \times 10^6$ million m^3

Chapter- 12
Air Pollution and Control

01. Ans: (c) **02. Ans: (d)**

03. Ans: (a) **04. Ans: (c)**

05. Ans: (d)

Sol: $CHCl_3 = 12 + 1 + 3\times35.5$

$= 119.5$ gm molecular.

Concentration = $0.4\ \mu g/m^3$

@ 273^0 k

$T_1 = 273^0$ k $T_2 = 293^0$k

$P_1 = 1$ $P_e = 1$

$V_1 = 22.4$ lit/ mol V_2 =?

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$\frac{22.4}{273} = \frac{V_2}{293} \Rightarrow V_2 = 24.04 \text{ lit/mol}$$

$$\mu g/m^3 = \frac{\text{PPm}\times\text{gm.mole.}\times10^3}{\text{lit/mol}}$$

$$0.4 = \frac{\text{PPm}\times119.5\times10^3}{24.04}$$

PPM = $8.047\times10^{-5} \times10^3$ $[\because 1\text{billion} = 10^9]$

$= 0.08$

Parts per billion

$= 8.047 \times10^{-5} \times10^3 = 0.08$

06. Ans: (b)

07. Ans: (a)

Sol: Dry air cools at 9.8^0 per km

$\approx 10^0$ km

For $\frac{1}{2}$km $\rightarrow 5^0$ fall

Final temperature at 500 m elevation

$= 40^0 - 5^0 = 35^0$c

08. Ans: (c)

Sol: Initial dry cot = 9.787 g, rate

= 1.5 m^3/min for 24 m^3

Final = 10.283 g

$$= \frac{(w_{\text{Initial}} - w_{\text{final}})}{\text{Volume of air sample}} \text{ filter paper}$$

Volume of are sample = rate of rate simplify × duration

= 145 × 24 × 60 = 2088 m^3

Total suspended particulate

$$= \frac{(10.283 - 9.787)}{2088}$$

= 2.375 × 10^{-4} gm/m^3

= 2.375 × 10^{-4} × 10^6

= 237.5 μg/m^3

09. Ans: (a) **10. Ans: (b)**

11. Ans: (d) **12. Ans: (c)**

13. Ans: (c)

14. Ans: (c)

Sol: 200001 m, No. = 50,000, rate

= 2 gm/km/vehicle

$$= 2 \times 50{,}000 \times 20000 \times \frac{1}{10^6}$$

= 2000 t

15. Ans: (b) **16. Ans: (a)**

17. Ans: (b) **18. Ans: (c)**

19. Ans: (a) **20. Ans: (c)**

21. Ans: 0.011268

Sol: $\frac{P}{RT} = 41.6\,\text{mo}\ell/\text{m}^3$

$$\frac{RT}{P} \times 10^3 = \text{constant} = 24.038$$

$$1\,\mu g/m^3 = \frac{24.038}{M \times 10^3}\,\text{ppm}$$

$$30\ \mu g/m^3 \text{ of } SO_2 = \frac{24.038}{64 \times 10^3} \times 30$$

= 0.011268 ppm

22. Ans: (b)

Sol: $v = \frac{2}{60}\,m/\text{sec}$

Total surface area of bags required $= \frac{Q}{v}$

$$\frac{10}{2/60} = 300\ m^2$$

Surface area of each bag = πdH

= π× 0.45 × 7.5

$$\text{Number of bags required} = \frac{300}{\pi \times 0.45 \times 7.5}$$

$= 28.29 \simeq 29$

23. Ans: 8012.38

Sol: η of ESP is given by

$$\eta = 1 - e^{\frac{-AW}{Q}}$$

when η = 96% A = 5600 m^2

η = 97% A = 6100 m^2

$\eta = 99\%$ $A = ?$

$0.96 = 1 - e^{\frac{-5600\times w}{185}} \Rightarrow W = 0.1063$ m/sec

$0.99 = 1 - e^{\frac{-A\times 0.1063}{185}} \Rightarrow A = 8012.38\ m^2$

plate area of collector $A = 8012.38\ m^2$

24. Ans: 55.55

Sol: Given coal consuming rate = 5 t/hr

Sulphur content = 2%

We know

$S + O_2 = SO_2$

32g 64 g

32 g produce 64 g of SO_2

1g of sulphur produce 2g of SO_2

Now sulphur consumption rate

= (0.02) × 5 t/hr

= $0.02 \times 5 \times 10^6$ gm/hr

$= \frac{2\times 5\times 10^4}{60\times 60}$ gm/sec

= 27.777 gm/sec

SO_2 emission rate = 2 × 27777

= 55.55 gm/cc

Chapter- 13
Noise Pollution

01. Ans: (b)

Sol: P_{rms}

$L_1 = 60\ dB = 20\log_{10}\frac{P_1}{P_0} = 20\log_{10}\frac{P_1}{20}$

$P_1 = P_2$

$P_{rms} = \sqrt{(20000)^2 + (20000)^2}$

= 28284.27 dB

Total SPL $= 20\log_{10}\frac{P_{rms}}{P_0}$

$= 20\log_{10}\frac{28284.27}{20} = 63$ dB (or)

(or) 60 + 3 = 63 dB

02. Ans: (b)

Sol: $L_1 = 80$ dB, $L_2 = 60$ dB

80 + 0 = 80 dB

03. Ans: (a)

Sol: $I = 20\log_{10}\times \frac{1}{4}\Sigma\left(10^{\frac{20}{20}} + 10^{\frac{56}{20}} + 10^{\frac{66}{20}} + 10^{\frac{42}{20}}\right)$

= 56.8 Db

04. Ans: 87.30

Sol: $L_{eq} = 10\log_{10}\left(\sum_{i=1}^{n} 10^{L_i/10}\times t_i\right)$

$L_{eq} = 10\log_{10}\left[10^{80/10}\times\frac{10}{95} + 10^{60/10}\times\frac{80}{95} + 10^{100/10}\times\frac{5}{95}\right]$

L_{eq}= 87.30 dB

05. Ans: (c)

Sol: P = 2000 μ bar

Particular = 10^5 N/m^{2}cPa)

1 μ bar $10^{-6} \times 10^5$ N/m^2 = 0.1 Pa

$\Rightarrow 10^{-6} \times 10^5 \times 10^6 = 10^5$ μPa

1 μ bar = 10^6 MPa

$1 = 20 \log_{10} \frac{2000\times10^5}{20} = 140$ dB

06. Ans: (b)

07. Ans: (a)

08. Ans: (a)

Sol: $20\log_{10}\frac{P}{Pref} = L$

P = 0.0002 μ

$L = 20\log_{10}\frac{2.0002\times10^5}{20}$

= 0 dB

When 2 (or) more source emited sound at a time then total sound pressure level

$= 20\log_{10}\frac{Prms}{P_0}$

Source $L_1 \to P_1$

Source $L_2 \to P_2$

Source $L_3 \to P_3$

$L_1 = 20\log_{10}\frac{P_1}{P_0}$

09. Ans: (c)

Sol: 50 dB + 50 dB

50 + 3 = 53 dB

10. Ans: 74 dB

Sol: L_1 = 80 dB, r_1 = 20 m

L_2 = ?, r_2= 40 m

$L_2 = L_1 - 20\log_{10}\left(\frac{r_2}{r_1}\right)$

$L_2 = 80 - 20\log_{10}\left(\frac{40}{20}\right)$

L_2 = 73.98 dB

$\simeq$ 74dB